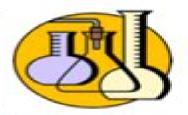
## Stoichiometry and The Mole – Review Guide

## Remember:

- 1 mole = 6.02 x 10<sup>23</sup> atoms, molecules, formula units
- 1 mole = grams (of molar mass)
- 1 mole = 22.4 liters (L) of gas



- The process used to predict the amount of a product that will be formed is \_\_\_\_\_\_stoichiometry
- 2) If you know the mass of a substance, you can find the number of moles by dividing it by the...

#### molar mass of substance

 If you know the <u>number of particles (atoms)</u> of a substance, you can find the number of moles by dividing it by...

## Avogadro's Number

4) Summarize the steps of a gram-to-gram stoichiometry problem:

Grams A → Moles A → Moles B → Grams B

The essential stoichiometry problem needed to convert from one chemical to another is:

Mole to Mole

6) Two reactants undergo a chemical reaction. At the end of the reaction, Chemical A still exists in the container, while Chemical B has been completely used up. Which is the limiting reactant?

#### Chemical B

- 7) What is a mole? \_\_\_\_ A unit of measurement used to count atoms/molecules.
- 8) What is Avogadro's Number? 6.02 x 10<sup>23</sup>
- 9) What does Avogadro's number tell us? <u>The number of atoms or molecules in a mole.</u>

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**Morton M. Sternheim** 

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