

## SHORT QUESTIONS

**Q.1 Aromatic compounds burn with sooty flame, why?**

**Ans.** Aromatic compounds have high percentage of carbon as compared to hydrogen. Aromatic compounds containing six carbon atoms have six hydrogen in it but the aliphatic compounds with six carbon atoms have fourteen hydrogen atoms aromatic compounds due to greater percentage of carbon, do not burn properly and give soot or smoke of carbon.

**Q.2 What is meant by aromatic compounds?**

**Ans.** The term aromatic was derived from the Greek word "aroma" meaning "fragrant" usually the compounds containing six membered benzene ring are called aromatic compound e.g., phenol, toluene, benzene, naphthalene, etc.

**Q.3 Benzene usually undergoes electrophilic substitution reaction and not addition reactions, why?**

**Ans.** In benzene ring, delocalized  $\pi$ -electrons are present in the form of sheath. This sheath of electrons above and below the carbon plane can be donated to the substance, which is electron deficient. During electrophilic substitution reactions, the aromaticity of ring remains as such. In addition reaction of benzene it changes to a non-aromatic compounds.

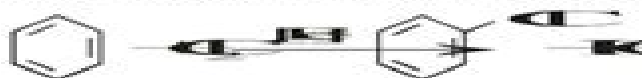


**Q.4 Benzene is less reactive than alkanes, why?**

**Ans.** In benzene ring, delocalized  $\pi$ -electrons are present in the form of sheath above and below the carbon plane. In case of ethene, localized  $\pi$ -electron cloud is present. Delocalized cloud of electrons is more stable as compared to localized cloud. Some energy is required to break the delocalization of electronic cloud called resonance energy. Due to this resonance stability benzene is less reactive than alkene.

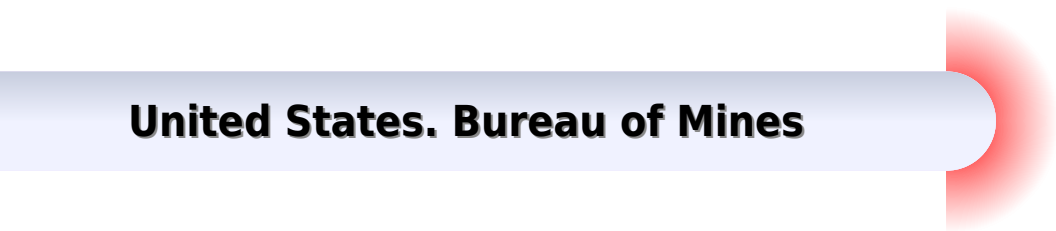
**Q.5 Give two reactions of benzene in which it behaves as saturated hydrocarbon.**

**Ans.** (i) Benzene reacts with chlorine in the presence of Fe to form chlorobenzene.



# Study Guide For Content Mastery Aromatic Hydrocarbons

**United States. Bureau of Mines**



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## **Table of Contents Study Guide For Content Mastery Aromatic Hydrocarbons**

1. Understanding the eBook Study Guide For Content Mastery Aromatic Hydrocarbons
  - The Rise of Digital Reading Study Guide For Content Mastery Aromatic Hydrocarbons
  - Advantages of eBooks Over Traditional Books
2. Identifying Study Guide For Content Mastery Aromatic Hydrocarbons
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Study Guide For Content Mastery Aromatic Hydrocarbons
  - User-Friendly Interface

4. Exploring eBook Recommendations from Study Guide For Content Mastery Aromatic Hydrocarbons
  - Personalized Recommendations
  - Study Guide For Content Mastery Aromatic Hydrocarbons User Reviews and Ratings
  - Study Guide For Content Mastery Aromatic Hydrocarbons and Bestseller Lists
5. Accessing Study Guide For Content Mastery Aromatic Hydrocarbons Free and Paid eBooks
  - Study Guide For Content Mastery Aromatic Hydrocarbons Public Domain eBooks
  - Study Guide For Content Mastery Aromatic Hydrocarbons eBook Subscription Services
  - Study Guide For Content Mastery Aromatic Hydrocarbons Budget-Friendly Options
6. Navigating Study Guide For Content Mastery Aromatic Hydrocarbons eBook Formats
  - ePub, PDF, MOBI, and More
  - Study Guide For Content Mastery Aromatic Hydrocarbons Compatibility with Devices
  - Study Guide For Content Mastery Aromatic Hydrocarbons Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Study Guide For Content Mastery Aromatic Hydrocarbons
  - Highlighting and Note-Taking Study Guide For Content Mastery Aromatic Hydrocarbons
  - Interactive Elements Study Guide For Content Mastery Aromatic Hydrocarbons
8. Staying Engaged with Study Guide For Content Mastery Aromatic Hydrocarbons
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Study Guide For Content Mastery Aromatic Hydrocarbons
9. Balancing eBooks and Physical Books Study Guide For Content Mastery Aromatic Hydrocarbons
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Study Guide For Content Mastery Aromatic Hydrocarbons
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Study Guide For Content Mastery Aromatic Hydrocarbons
  - Setting Reading Goals Study Guide For Content Mastery Aromatic Hydrocarbons
  - Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Study Guide For Content Mastery Aromatic Hydrocarbons
  - Fact-Checking eBook Content of Study Guide For Content Mastery Aromatic Hydrocarbons
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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