OMPTER (2) STUDY GUIDE FOR CONTENT MASTERS

Section 12.3 Limiting Reactants

in your textbook, read about any mactions stop and how to determine the Uniting reaction.

Study the diagram showing a chemical reaction and the elemical equation that represents the traction. Then complete the table. Show your calculations for questions 25-27 in the space below the table.

$$\Box + 00 + 0 \rightarrow 0 + 00$$

 $0,+200 \rightarrow 200$

The media masses of O₂, NO, and NO₂ are 32.00 g/mel, 30.01 g/mel, and 46.01 g/mel, emporatedly.

Amount of O ₂	Annual of NO	Amount of Mily	Similing Number	Amount and Name of Quees Resident
1 molecule	Teolesies	2 molecules	mine	none
Challetales	4 mileciles	4 molecules	NO	2 milesia 0 _y
I molecules	1 militaire	1. 4 molecules	2. 02	1, 4 molecules NO
3.88 mol -	2.00 mail	4. 2.00 mol	5. none	6. none
438 mol	450 mail	7, 4.00 mail	E. NO	9, 2.00 mol 0 ₂
3.00 mol	7.00 mail	16, 7.00 mol	11, NO	12, 1.50 mol O ₂
1.00 mol	400 mil	13. 2.00 mol	14. O ₃	15. 2.00 mol NO
1.500 mail	0.200 mail	16, 0,200 mol	17, 100	18. 0.400 mol O ₂
name	mag	19. 92.02 g	20. none	21. none
26.00 p.	mmy	22, 46.01 g	20. 02	34. 50.12 g NO
mm ₁	20.00 g	25, 28.7% g	26. 02	37, 1,34 g NO

balanced equation mole ratio = 2 mol NO/1 mol O2

10.00 g.O₂ × 1 mol O₂/32.00 g.O₂ = 0.3125 mol O₂

20.00 p.NO × 1 mai NO/30.01 p.NO = 0.6664 mai NO

actual mole ratio = 0.6664 mol NO/0.3125 mol O₂ = 2.132 mol NO/1.000 mol O₂ Secause the actual mole ratio of NO:O₂ is larger than the balanced equation mole ratio of NO:O₃, there is an excess of NO; O₃ is the limiting reactant.

Mass of NO used = 0.3125 gpai-0; × 2 mol NO¹1 gpai-0; = 0.6250 mol NO 0.6250 gpai.NO × 30.01 g NO¹1 gpai-400 = 18.76 g NO

Mass of NO₂ produced = 0.6250 gpsi.WO₂ × 46.01 g NO₂/1 gpsi.WO₂ = 28.76 g NO₂ Excess NO = 20.00 g NO - 18.76 g NO = 1.24 g NO

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Chemistry Matter and Change + Chapter 12 ... 29.



Section 12.4 Percent Yield

In your textlisiol, resel about the pielos of products.

Study the diagram and the example problem.

mass of product from experimental measurement

mais of product predicted from assistionietric calculation using

- at mass of reactions.
- 5. 4-dep main to main conversion
 - 1. Sems the belanced chemical equation.
 - Calculate the number of moles of reactions, using moles mass.
 - Calculate the number of moles of product, using the appropriate mole ratio.
 - Calculate the mass of product, using the reciprocal of molar mass.

Example Problem: The following chemical equation represents the production of gallium oriole, a substance used in the manufacturing of some semiconductor devices.

In one experiment, the reaction yielded 7.42 g of the enals from a 7.85 g sample of gallium. Describes the percent yield of this reaction. The molar masses of Ga and Ga₂O₂ are 46.72 glossl and 167.46 glossl, respectively.

Use the information in the diagram and example problem to evaluate each value or expression below. If the value or expression is correct, write correct. If it is incorrect, write the correct value or expression.

- 1. actual right unknown 7.42 g Ga₂O₃
- 2. mass of touchast: Total y the connect.
- matter of motion of reactant: 7.00 g/G₀ × min(T₂ g/G₀).
 mod G₀.
 mod G₀.
 mod G₀.
 mod G₀.
- 4. number of moles of product 0.000 mol $G_0 \times \frac{2 \bmod G_0/O_0}{1 \bmod G_0} = \frac{0.100 \bmod G_0 \times 2 \bmod G_0}{2 \bmod G_0/O_0/4 \bmod G_0}$
- 5. theoretical yield. 0:0500 and $O_{m_0}O_{\gamma} \times \frac{107.44 \pm O_{m_0}O_{\gamma}}{1 \text{ and } O_{m_0}O_{\gamma}}$ connect.
- percent yield: 41.7 g Ga₂O₃ × 100
 7.42 g Ga₂O₃/9.37 g Ga₂O₃ × 100
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Jon Wiles, Joseph Bondi

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Handbook on Measurement, Assessment, and Evaluation in Higher Education Charles Secolsky, D. Brian Denison, 2017-07-31 In this valuable resource well known scholars present a detailed understanding of contemporary theories

and practices in the fields of measurement assessment and evaluation with guidance on how to apply these ideas for the benefit of students and institutions Bringing together terminology analytical perspectives and methodological advances this second edition facilitates informed decision making while connecting the latest thinking in these methodological areas with actual practice in higher education This research handbook provides higher education administrators student affairs personnel institutional researchers and faculty with an integrated volume of theory method and application **Teaching, and Learning** National Research Council, Division of Behavioral and Social Sciences and Education, Board on Testing and Assessment, Committee on Title I Testing and Assessment, 1999-10-06 State education departments and school districts face an important challenge in implementing a new law that requires disadvantaged students to be held to the same standards as other students The new requirements come from provisions of the 1994 reauthorization of Title I the largest federal effort in precollegiate education which provides aid to level the field for disadvantaged students Testing Teaching and Learning is written to help states and school districts comply with the new law offering guidance for designing and implementing assessment and accountability systems This book examines standards based education reform and reviews the research on student assessment focusing on the needs of disadvantaged students covered by Title I With examples of states and districts that have track records in new systems the committee develops a practical decision framework for education officials The book explores how best to design assessment and accountability systems that support high levels of student learning and to work toward continuous improvement Testing Teaching and Learning will be an important tool for all involved in educating disadvantaged students state and local administrators and classroom teachers Unstoppable Learning Yazmin Pineda Zapata, Rebecca Brooks, 2017-05-30 Foreword by Douglas Fisher and Nancy Frey Modify your curriculum using streamlined assignments targeted assessments and student engagement strategies for different learning styles This practical guide expands upon the Unstoppable Learning model to explore accessible learning for all students Through specific curriculum and environmental accommodations and modifications as well as personal and technology supports K 12 teachers will discover how to provide differentiated instruction to students with varying needs from physical disabilities to twice exceptionality Forms tools and diagrams designed to aid instructional planning are also included How will this book help you Learn how to adjust curriculum in ways that maintain appropriate levels of rigor for different learning styles Consider real school vignettes and examples that illustrate successful Unstoppable Learning adaptations for inclusive classrooms Study the four guiding principles of systems thinking relationships communications responsiveness and sustainability Consider the importance of collaboration and communication in learning adaptations Study helpful and informative tools and diagrams intended to aid in planning instruction for different learning styles Teaching at Its Best Linda B. Nilson, 2010-04-20 Teaching at Its Best This third edition of the best selling handbook offers faculty at all levels an essential toolbox of hundreds of practical teaching techniques formats classroom activities and exercises all of which can be

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