

CHAPTER 6

STUDY GUIDE

Ionic Compounds and Metals

Lesson 6.1 Ion Formation

In your textbook, read about chemical bonds and formation of ions.

Use each of the terms below just once to complete the passage.

chemical bond	electrons	energy level	ions	noble gases
nucleus	octet	pseudo-noble gas formations	valence	

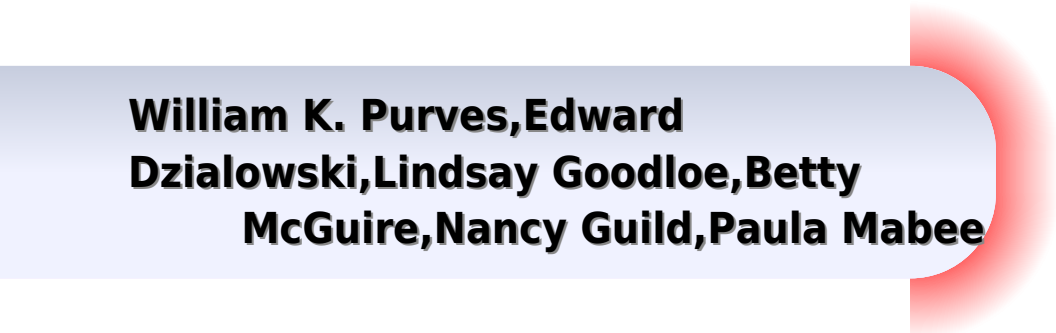
The force that holds two atoms together is called a(n) (1) Chemical Bond. Such an attachment may form by the attraction of the positively charged (2) Nucleus of one atom for the negatively charged (3) Electron of another atom, or by the attraction of charged atoms, which are called (4) Ions. The attractions may also involve (5) Valence electrons, which are the electrons in the outermost (6) Energy level. The (7) Noble Gas are a family of elements that have very little tendency to react. Most of these elements have a set of eight outermost electrons, which is called a stable (8) Octet. The relatively stable electron structures developed by loss of electrons in certain elements of groups 3, 4, 13, and 14 are called (9) Pseudo-Noble gas formations.

For each statement below, write *true* or *false*.

- | | |
|--------------|--|
| <u>False</u> | 10. A positively charged ion is called an anion. |
| <u>True</u> | 11. Elements in group 1 lose their one valence electron, forming an ion with a 1+ charge. |
| <u>False</u> | 12. Elements tend to react so that they acquire the electron structure of a halogen. |
| <u>True</u> | 13. A sodium atom tends to lose one electron when it reacts. |
| <u>True</u> | 14. The electron structure of a zinc ion (Zn^{2+}) is an example of a pseudo-noble gas formation. |
| <u>False</u> | 15. A Cl^- ion is an example of a cation. |
| <u>True</u> | 16. The ending <i>-ide</i> is used to designate an anion. |
| <u>False</u> | 17. Nonmetals form a stable outer electron configuration by losing electrons and becoming anions. |

Study Guide For Ions And Ionic Compounds

**William K. Purves, Edward
Dzialowski, Lindsay Goodloe, Betty
McGuire, Nancy Guild, Paula Mabee**



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Study Guide and Solutions Manual for Fundamentals of General, Organic, and Biological Chemistry Susan McMurry, 1992

The Practice of Chemistry Study Guide & Solutions Manual Pamela Mills, Amina El-Ashmawy, 2003-04-14 Designed to help students understand the material better and avoid common mistakes Also includes solutions and explanations to odd numbered exercises **Study Material Based On NCERT Science Class- X** Dr. Sunita Bhagiya, ,Er. Meera

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Organic Chemistry, 12e Binder Ready Version Study Guide & Student Solutions Manual T. W. Graham Solomons, Craig B. Fryhle, Scott A. Snyder, 2016-04-11 This is the Student Study Guide Solutions Manual to accompany Organic Chemistry 12th Edition The 12th edition of Organic Chemistry continues Solomons Fryhle Snyder's tradition of excellence in teaching and preparing students for success in the organic classroom and beyond A central theme of the authors' approach to organic chemistry is to emphasize the relationship between structure and reactivity To accomplish this the content is organized in a way that combines the most useful features of a functional group approach with one largely based on reaction mechanisms The authors' philosophy is to emphasize mechanisms and their common aspects as often as possible and at the same time use the unifying features of functional groups as the basis for most chapters The structural aspects of the authors' approach show students what organic chemistry is Mechanistic aspects of their approach show students how it works And wherever an opportunity arises the authors show students what it does in living systems and the physical world around us Study Material Based On NCERT Science Class - IX Dr. Sunita Bhagiya, ,Er. Meera

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