

FIRST TERM EXAMINATION 2022-23

: 12th (Sci.)

Subject : Chemistry

Marks : 50

Time : 2:30 hrs.

- 1) Question paper are in four section containing 23 questions. 2) Use of long table is allowed.
3) For each MCQ correct answer must be written along with its alphabet.

Section : A

Select and write the correct answer.

(7)

Which of the following is n-type semiconductor.

- a) Pure Si b) Si doped with As c) Si doped With Ga d) Ge doped with in

Cryoscopic constant depends on.

- a) Nature of solvent b) Nature of solute c) Nature of solution

d) Number of solvent molecules.

The solubility product of a sparingly soluble salt AX is 5.2×10^{-10} . It is solubility in mol dm^{-3} is

- a) 7.2×10^{-7} b) 1.35×10^{-4} c) 7.2×10^{-4} d) 13.5×10^{-4}

Which of the following are not state function? In alides halides, the halogen

- 1) Q+W 2) Q 3) W 4) H-TS atom is bonded 2 Hybridize Carbon

a) 1, 2 and 3 b) 2 and 3 c) 1 and 4 d) 2, 3 and 4 a) S.P. b) S.P. c) S.P. d) None of them

SI unit of molar conductivity (Λ)

- a) $\text{S cm}^2 \text{ mol}^{-1}$ b) $\text{S dm}^2 \text{ mol}^{-1}$ c) S m^2 d) $\text{S m}^2 \text{ Mol}^{-1}$

The elementary reaction, $2\text{NO} + \text{Cl}_2 \longrightarrow 2\text{NOCl}$ is rate = $k[\text{NO}]^2[\text{Cl}_2]$. This of K would increase with

- a) Increase of temperature b) increase of concentration of NO
c) Increase of Concentration of Cl_2 d) increase of concentration of both Cl_2 and NO

BrCl reacts with water to form allikes halides - - - - compounds a) polar

- a) HBr b) $\text{Br}_2 + \text{Cl}_2$ c) HOBr d) $\text{HOBr} + \text{HCl}$ b) non-polar

Answer the following.

Calculate van't Hoff's factor assuming complete dissociation of potassium bromide solution. a) Soluble b) None of them

Define Entropy. Solution

Write electrode reaction during electrolysis of molten KCl.

Define hydrolysis of salt.

Define Isomorphism.

What are inter-halogen compound? What is order of reaction.

Calculate the spin only magnetic moment of divalent cation of a transition metal with atomic

number 25. Calculate the molar conductivity of KCl at molar ionic

SECTION B

Attempt any eight

Distinguish between Lanthanoids and Actinoids.

Elements of group 16 generally show lower values of first ionisation enthalpy compared to

the elements of corresponding period of group 15 why?

Why alkyl halide insoluble in water explain

(16)

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Rose Arny



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