

Unit 9 Worksheet: St. Michael's Concentration

Part 9: Describing Molar Concentrations

1. Sodium chloride was dissolved in water to produce a 1.5M solution. Suppose that this same solution had been added to the flask solution. This would be incorrect. **Just this solution.**
Then, calculate the molarities of the two solutions. The molarity of the solution is 1.5M. The molarity of the solution is 1.5M.
2. A 10.0 g solution of potassium chloride is dissolved in enough water to make 100.0 mL of solution. Calculate the molar concentration of this potassium chloride.

$$10.0 \text{ g} = \frac{1 \text{ mol}}{74.55 \text{ g/mol}} \times 100.0 \text{ mL}$$

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3. Find the molarity of a solution made from 25.0 g of CaCl_2 dissolved in enough water to make 0.5 L.

$$25.0 \text{ g} = \frac{1 \text{ mol}}{110.98 \text{ g/mol}} \times 0.5 \text{ L}$$

$$25.0 \text{ g} = \frac{1 \text{ mol}}{110.98 \text{ g/mol}} \times 0.5 \text{ L}$$

4. An unknown liquid solution of "molarities" is dissolved in enough water to make 0.5 L of solution. Suppose the molarity of the solution is 1.5M. Calculate the molarity of the solution in the flask.

$$0.5 \text{ L} = \frac{1 \text{ mol}}{1.5 \text{ M}} \times 100.0 \text{ mL}$$

$$0.5 \text{ L} = \frac{1 \text{ mol}}{1.5 \text{ M}} \times 100.0 \text{ mL}$$

5. Suppose that a solution of 1.5M is dissolved in enough water to make 0.5 L of solution. Calculate the molarity of the solution in the flask.

$$1.5 \text{ M} = \frac{1 \text{ mol}}{1.5 \text{ M}} \times 0.5 \text{ L} = 0.5 \text{ L} \times 1.5 \text{ M} = 0.75 \text{ M}$$

Unit 1 worksheet 3 Molar Concentration

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