

Name: _____

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UNIT 2 – BIOCHEMISTRY TEST STUDY GUIDE

Basic Chemistry

Use the terms below to complete the statements below.

(HINT: Some of the words will be used in multiple statements.)

Electrons
Electron cloud
Ion

Isotopes
Neutral
Neutrons

Nucleus
Proton

The two regions of the Atom are the nucleus and the electron cloud.

The particles found within an atom are the proton, neutron and the electron.

The mass number of an element is equal to the protons and neutrons. The mass of an atom is located in the nucleus.

Since ATOMS are neutral the protons and the electrons will always be equal. If they weren't, then there would be a charge and the atom would be called an ion.

Atoms that have the same number of protons, but have different numbers of neutrons are called isotope.

Complete the table below describing the particles in the atom

Particle	Abbreviation	Location	Charge
Proton	P ⁺	Nucleus	positive
Neutron	N ⁰	Neucleus	No charge
Electron	E ⁻	Electron cloud	Negative

Label the periodic table block below with the following terms:

element name, element symbol, average atomic mass, & atomic number.

30	Zn
	Zinc
	65.39

What information does the atomic number tell you about the atom?

The number of protons in the atom.

What information does a mass number tell you about the atom?

The number of protons and neutrons in the atom

How can you calculate the number of neutrons in an atom? Atomic mass – atomic number

List the names and symbols of the elements described below:

Elements that make up 96% of living matter

Element Name Carbon	Element Symbol C	Element Name Hydrogen	Element Symbol H	Element Name Oxygen	Element Symbol O	Element Name Nitrogen	Element Symbol N
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What are the other 4% of elements called? Minerals and Trace elements

Unit 2 Biochemistry Study Guide

Barbara R. Sandall, Ed.D.



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