

AP Chemistry Unit 4: Chemical Reactions
Worksheet 1

- Calculate the mass percent of oxygen in each of the following compounds:
(a) CuCO_3 (b) $\text{Fe}(\text{HCO}_3)_2$
- Calculate the mass of oxygen in each of the following samples:
(a) 2.45 g of $\text{C}_4\text{H}_8\text{O}_3$ (b) 7.10 g of $\text{Fe}(\text{HCO}_3)_2$
- A certain compound has a formula of Cu_xO_3 , where x is the subscript on the copper ion. The mass percent of the oxygen in a 2.0 g sample is found to be 11.1%.
(a) How many moles of oxygen are present in the sample?
(b) What is the molar mass of the compound?
(c) What is the subscript on the Cu?
$$6 \text{Li}(s) + \text{N}_2(g) \rightarrow 2 \text{Li}_3\text{N}(s)$$
- If 1.50 g of lithium reacts full on above the reaction:
(a) How many grams of N_2 also react?
(b) How many grams of Li_3N are created?
$$\text{MgCO}_3(s) + 2 \text{HCl}(aq) \rightarrow \text{MgCl}_2(aq) + \text{CO}_2(g) + \text{H}_2\text{O}(l)$$
- When the above reaction goes to completion, 245 mL of CO_2 is produced at STP.
(a) What mass of CO_2 is present?
(b) What mass of MgCO_3 reacted?
(c) How many water molecules were produced?
$$2 \text{NaNH}_2(s) \rightarrow 2 \text{Na}(s) + \text{N}_2(g)$$
- 3.78×10^{22} molecules of N_2 are created when some NaNH_2 decomposes.
(a) How many moles of NaNH_2 decomposed?
(b) How many moles of Na are created?

Unit 4 Section Quiz A Chemistry

ML Morrison



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Table of Contents Unit 4 Section Quiz A Chemistry

1. Understanding the eBook Unit 4 Section Quiz A Chemistry
 - The Rise of Digital Reading Unit 4 Section Quiz A Chemistry
 - Advantages of eBooks Over Traditional Books
2. Identifying Unit 4 Section Quiz A Chemistry
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Unit 4 Section Quiz A Chemistry
 - User-Friendly Interface
4. Exploring eBook Recommendations from Unit 4 Section Quiz A Chemistry
 - Personalized Recommendations

- Unit 4 Section Quiz A Chemistry User Reviews and Ratings
- Unit 4 Section Quiz A Chemistry and Bestseller Lists
- 5. Accessing Unit 4 Section Quiz A Chemistry Free and Paid eBooks
 - Unit 4 Section Quiz A Chemistry Public Domain eBooks
 - Unit 4 Section Quiz A Chemistry eBook Subscription Services
 - Unit 4 Section Quiz A Chemistry Budget-Friendly Options
- 6. Navigating Unit 4 Section Quiz A Chemistry eBook Formats
 - ePub, PDF, MOBI, and More
 - Unit 4 Section Quiz A Chemistry Compatibility with Devices
 - Unit 4 Section Quiz A Chemistry Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Unit 4 Section Quiz A Chemistry
 - Highlighting and Note-Taking Unit 4 Section Quiz A Chemistry
 - Interactive Elements Unit 4 Section Quiz A Chemistry
- 8. Staying Engaged with Unit 4 Section Quiz A Chemistry
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Unit 4 Section Quiz A Chemistry
- 9. Balancing eBooks and Physical Books Unit 4 Section Quiz A Chemistry
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Unit 4 Section Quiz A Chemistry
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Unit 4 Section Quiz A Chemistry
 - Setting Reading Goals Unit 4 Section Quiz A Chemistry
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Unit 4 Section Quiz A Chemistry
 - Fact-Checking eBook Content of Unit 4 Section Quiz A Chemistry

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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