

Unit 8 Worksheet 1: Mole relationships

For each of the problems below:

- Write the balanced chemical equation.
- Identify what is given (with units) and what you want to find (with units).
- Use coefficients from balanced equation to determine mole ratio.
- Show set up (organize it!).

- Hydrogen sulfide gas, which smells like rotten eggs, burns in air to produce sulfur dioxide and water. How many moles of oxygen gas would be needed to completely burn 8 moles of hydrogen sulfide?

Equation:	$\underline{\quad 2 \quad} \text{H}_2\text{S}_{(g)} + \underline{\quad 3 \quad} \text{O}_{2(g)} \rightarrow \underline{\quad 2 \quad} \text{SO}_{2(g)} + \underline{\quad 2 \quad} \text{H}_2\text{O}_{(g)}$				
Before	<u>2</u>	ex	<u>0</u>	<u>0</u>	3.0 mol O ₂
Change	<u>-2</u>	<u>-3</u>	<u>2</u>	<u>2</u>	
After	<u>0</u>	ex	<u>2</u>	<u>2</u>	

- Propane, C₃H₈, burns in air to form carbon dioxide and water. If 12 moles of carbon dioxide are formed, how many moles of propane were burned?

Equation: $2\text{AL} + 6\text{HCL} \rightarrow 2\text{ALCL}_3 + 3\text{H}_2$

Before	0.5	ex	0	0	1.50mol HCL
Change	-0.5	-1.5	+0.5	+0.5	
After	0	ex	0.5	0.75	

- Ammonia, NH₃, for fertilizer is made by causing hydrogen and nitrogen to react at high temperature and pressure. How many moles of ammonia can be made from 0.15 moles of nitrogen gas?

Equation: $3\text{H}_2 + \text{N}_2 \rightarrow 2\text{NH}_3$

Before	ex	1.25	0	2.50 mole
Change	-3.75	-1.25	+25	
After	ex	0	2.5	

Unit 8 Ws 1 Mole Relationships Answers

Dr. S C Rastogi,, Er. Meera Goyal

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the authors' names.

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Matter Gases and Liquids Unit VI Chemical Thermodynamics Unit VII Equilibrium Unit VIII Redox Reactions Unit IX Hydrogen Unit X s Block Elements Alkali and Alkaline earth metals Group 1 and Group 2 Elements Unit XI Some p Block Elements General Introduction to p Block Elements Unit XII Organic Chemistry Some Basic Principles and Techniques Unit XIII Hydrocarbons Classification of Hydrocarbons Unit XI V Environmental Chemistry Content 1 Some Basic Concepts of Chemistry 2 Structure of Atom 3 Classification of Elements and Periodicity in Properties 4 Chemical Bonding and Molecular Structure 5 States of Matter 6 Thermodynamics 7 Equilibrium 8 Redox Reactions 9 Hydrogen 10 s Block Elements 11 p Block Elements 12 Organic Chemistry Some Basic Principles and Techniques 13 Hydrocarbons 14 Environmental Chemistry I Appendix II Log antilog Table

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Solution Interface Precipitation and Dissolution Particle Particle Interaction Oxidation and Reduction Regulation of the Chemical Equilibria and Microbial Mediation Composition of Natural Waters Chemistry International ,1996

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Table of Contents Unit 8 Ws 1 Mole Relationships Answers

1. Understanding the eBook Unit 8 Ws 1 Mole Relationships Answers
 - The Rise of Digital Reading Unit 8 Ws 1 Mole Relationships Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Unit 8 Ws 1 Mole Relationships Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Unit 8 Ws 1 Mole Relationships Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Unit 8 Ws 1 Mole Relationships Answers
 - Personalized Recommendations
 - Unit 8 Ws 1 Mole Relationships Answers User Reviews and Ratings

- Unit 8 Ws 1 Mole Relationships Answers and Bestseller Lists
- 5. Accessing Unit 8 Ws 1 Mole Relationships Answers Free and Paid eBooks
 - Unit 8 Ws 1 Mole Relationships Answers Public Domain eBooks
 - Unit 8 Ws 1 Mole Relationships Answers eBook Subscription Services
 - Unit 8 Ws 1 Mole Relationships Answers Budget-Friendly Options
- 6. Navigating Unit 8 Ws 1 Mole Relationships Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Unit 8 Ws 1 Mole Relationships Answers Compatibility with Devices
 - Unit 8 Ws 1 Mole Relationships Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Unit 8 Ws 1 Mole Relationships Answers
 - Highlighting and Note-Taking Unit 8 Ws 1 Mole Relationships Answers
 - Interactive Elements Unit 8 Ws 1 Mole Relationships Answers
- 8. Staying Engaged with Unit 8 Ws 1 Mole Relationships Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Unit 8 Ws 1 Mole Relationships Answers
- 9. Balancing eBooks and Physical Books Unit 8 Ws 1 Mole Relationships Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Unit 8 Ws 1 Mole Relationships Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Unit 8 Ws 1 Mole Relationships Answers
 - Setting Reading Goals Unit 8 Ws 1 Mole Relationships Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Unit 8 Ws 1 Mole Relationships Answers
 - Fact-Checking eBook Content of Unit 8 Ws 1 Mole Relationships Answers
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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