

Unit 3 Review: Thermochemistry

A. Definition Questions: you should be able to use or apply the following terms. Be able to write complete definitions for the terms in **bold**.

energy

potential energy

kinetic energy

enthalpy

thermal kinetic energy

temperature

heat

specific heat capacity

endothermic

exothermic

standard state

allotropes

standard heat of formation

bond energy

1. Identify the **INCORRECT** statement below:

- a) potential energy is the energy an object has because of its position relative to and attraction toward another object.
- b) energy is the ability to do work.
- c) in an exothermic reaction, the enthalpy of the products is higher than the enthalpy of the reactants.
- d) in an ordinary chemical reaction, energy is neither created nor destroyed.

2. A particular chemical reaction has $\Delta H^\circ = +250 \text{ kJ/mol}$. Which of the following statements is/are true concerning this reaction?

I) Heat moves from the system to the surroundings

II) The reaction is endothermic

III) The enthalpy of the products is lower than the enthalpy of the reactants

- a) I and II only
- b) II only

- c) II and III only
- d) I, II and III are true

3. Which of the following molecules has the highest enthalpy?

- a) propane
- b) cyclopropane

- c) pentane
- d) cyclooctane

4. Which statement is **INCORRECT** about endothermic reactions?

- a) the system absorbs energy from its surroundings
- b) the enthalpy of products is lower than the enthalpy of the reactants
- c) the thermal kinetic energy of the surroundings will decrease
- d) the enthalpy change will have a positive value

5. All of the following statements are true **EXCEPT**:

- a) ΔH for a reaction is equal in magnitude but opposite in sign to ΔH for the reverse reaction
- b) enthalpy is a state function
- c) the overall change in enthalpy for a reaction depends on the number of steps in a reaction
- d) chemical potential energy contributes to the total enthalpy of a substance

6. What is the definition of the temperature of a substance?

- a) the total heat content of the particles in a substance
- b) the speed of the fastest particles in the substance
- c) the speed of the slowest particles in the substance
- d) the average kinetic energy of the particles in a system

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