Thermochemistry Review Part 1

Definition Questions: you should be able to use or apply the following terms. Be able to **write** complete definitions for,

energy	temperature	standard state
potential energy	heat	standard heat of formation
kinetic energy	specific heat capacity	exothermic
enthalpy	endothermic	Law of conservation of Energy
thermal	Hess's Law	isolated, closed, open system

1. The specific heat capacity of diamond is 0.5050 J/g°C. How much energy is required to heat 25.0 g of diamond from 10.5°C to 15.6°C?

$q = m \times c \times \Delta t$
$= 25\ g \times 0.5050\ J/g°C \times (15.6 - 10.5°C) = 64.39\ J$

2. A piece of aluminum with a mass of 50g and an initial temperature of 90°C is placed into 100mL of water at a temperature of 25°C. The temperature of water rises to 31.3°C. Determine the specific heat capacity of aluminum.

Q lost (Al) = Q gained (Water)
$-(mc\Delta t) = +mc\ \Delta t$
$-(50 \times c \times (31.3-90)) = (100 \times 4.18 \times (31.3-25))$
$-(-2935\ c) = 2633.4$
$c = 2633.4 / 2935 = 0.897\ J/g°C$

3. How much energy is absorbed by 300g of methanol, CH_3OH, as it evaporates? Given: The molar heat of vaporization is 35.3kJ/mol.

q = ?	$\Delta H = q/n$	$n\ CH_3OH = m/MM$
m = 300 g	$q = H \times n$	$n = 300g / 32\ g/mol$
$\Delta H = 35.3\ kJ/mol$	$q = (35.5\ kJ \times 9.375\ mol)$	$n = 9.375\ mol$
	$q = 332.81\ kJ$	

4. What is the molar enthalpy of the formation of 1mol $H_2SO_{4(l)}$ given the following information?

$H_2O_{(g)} + SO_{3(g)} \rightarrow H_2SO_{4(l)}$ $\quad \Delta H_{rxn} = x kJ$

Flip and divide by 2	$2H_{2(g)} + O_{2(g)} \rightarrow 2H_2O_{(g)}$	$\Delta H_{rxn} = -484kJ = +242$
Flip and divide by 2	$S_{(s)} + 3O_{2(g)} \rightarrow 2SO_{3(g)}$	$\Delta H_{rxn} = -890kJ = +445$
	$H_{2(g)} + 2O_{2(g)} + S_{(s)} \rightarrow H_2SO_{4(l)}$	$\Delta H_{rxn} = -814kJ = -814$
		$\Delta H_{rxn} = -127\ kJ$

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TD Snyder

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