

Unit 16 Test Review WS

Chemist: Elham Ghiassi

- The transfer of energy is called heat.
- The way we measure heat energy is temperature.
-



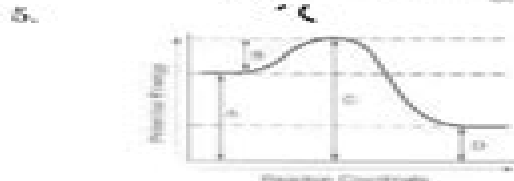
During an exothermic reaction, energy is released. The enthalpy or ΔH is negative. The reaction container will feel warmer. The enthalpy or energy of the product is lower than the enthalpy or energy of the reactants. If energy is shown in the chemical equation, it will be a products.

- Label each of the following as representing an endothermic or exothermic reaction:

Endothermic

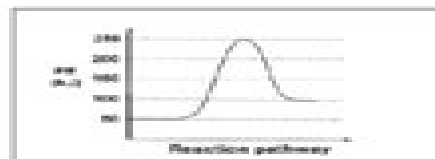


Exothermic

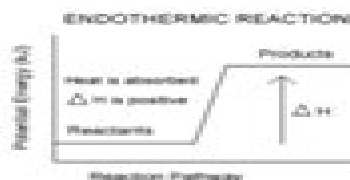


D minus A represents the ΔH for the forward reaction in the diagram to the left.

- Using the reaction curve to the right, calculate the enthalpy for the reaction. 50 kJ

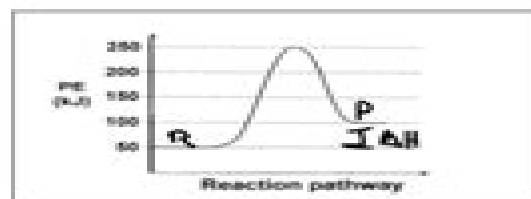


- If the enthalpy is negative, the reaction is exothermic (*endothermic/exothermic*).
- $N_2 + 2O_2 + \text{energy} \rightarrow 2NO_2$ Is this reaction endo- or exothermic? endothermic
-



In an endothermic reaction, heat is absorbed. The enthalpy is positive. The energy of the products is more than the energy of the reactants.

- Label the following on the graph to the right:
energy of reactants, energy of products,
enthalpy of the reaction



- The energy required to form the transition state (activated complex) from reactants is called the activation energy.

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