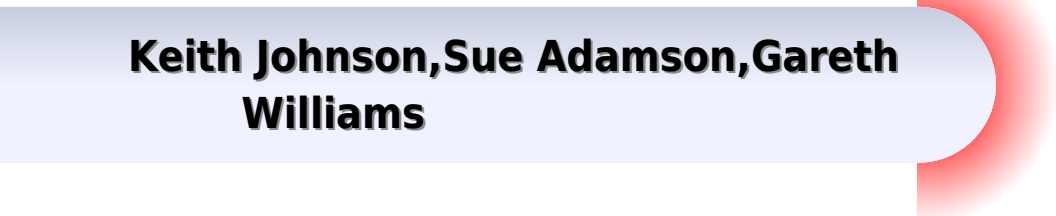


Unit 7 Chemical Reactions
PRACTICE TEST/ REVIEW

- ____ 1. A chemical reaction is a process in which
- a. products change into reactants
 - b. the law of conservation of mass applies
 - c. substances can change state
 - d. all of these
- ____ 2. During a chemical reaction,
- a. new elements are produced
 - b. atoms are rearranged
 - c. atoms are destroyed
 - d. elements are destroyed
- ____ 3. An equation is balanced by
- a. changing subscripts
 - b. adding coefficients
 - c. erasing elements as necessary
 - d. adding elements as necessary
- ____ 4. What are the reactants in the following chemical equation:
- $$\text{Zn} + \text{CuSO}_4 \rightarrow \text{ZnSO}_4 + \text{Cu}$$
- a. zinc and copper
 - b. zinc sulfate and copper
 - c. zinc and copper (II) sulfate
 - d. only zinc
- ____ 5. What are the products in the above equation?
- a. zinc and copper
 - b. zinc sulfate and copper
 - c. zinc and copper (II) sulfate
 - d. only zinc
- ____ 6. The breaking down of a substance into two or simpler substances is _____
- a. Displacement
 - b. Synthesis
 - c. Decomposition
 - d. Double displacement
- ____ 7. Each substance to the right of the arrow in a chemical equation is a(n)
- a. Reactant
 - b. Catalyst
 - c. Product
 - d. Inhibitor
- ____ 8. According to the law of conservation of mass, if two atoms of hydrogen are used as a reactant, how many atoms of hydrogen must be part of the product?
- a. 4
 - b. 2
 - c. 1
 - d. 0
- ____ 9. What type of reaction is shown in the following chemical equation, $\text{NH}_4\text{Cl} \rightarrow \text{NH}_3 + \text{HCl}$?
- a. Displacement
 - b. Synthesis
 - c. Decomposition
 - d. Double displacement

Unit 7 Chemical Reactions Test

**Keith Johnson, Sue Adamson, Gareth
Williams**



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