

# Gas Laws HOMEWORK

## Lesson 1 Homework

1. Define pressure. *Force that a gas exerts on its container*

2. Convert the following:

a. 750 mmHg to atm

b. 550 mmHg to atm

c. 2.5 atm to torr

d.  $2.1 \times 10^3$  Pa to atm

e. 300 Pa to atm

$$750 \text{ mmHg} \times (1 \text{ atm} / 760 \text{ mmHg}) = 0.9868 \text{ atm}$$

$$550 \text{ mmHg} \times (1 \text{ atm} / 760 \text{ mmHg}) = 0.72368 \text{ atm}$$

$$2.5 \text{ atm} \times (760 \text{ torr} / 1 \text{ atm}) = 1900 \text{ torr}$$

$$2.1 \times 10^3 \text{ Pa} \times (1 \text{ atm} / 101325 \text{ Pa}) = 0.02072 \text{ atm}$$

$$300 \text{ Pa} \times (1 \text{ atm} / 101325 \text{ Pa}) = 2.96 \times 10^{-3} \text{ atm}$$

3. Calculate the atmospheric pressure in kPa on a day when the pressure is 750 mm.

$$750 \text{ mm} \times (101.325 \text{ kPa} / 760 \text{ mm}) = 100.367 \text{ kPa} \rightarrow 100.4 \text{ kPa}$$

4. Calculate the pressure of a sample of gas (in mmHg) inside an open-end manometer if the mercury column is 35 mm higher on the side of the closed tube than on the atmosphere. The atmospheric pressure is 750 mm. Draw a picture to illustrate your answer.

$$750 \text{ mm} + (35 \text{ mmHg} / 1 \text{ mm}) = 785 \text{ mmHg}$$

*750 mm mercury is 750 mm Hg because the atmospheric pressure, then it's 750 mmHg because 750 mmHg.*

*Since pressure doesn't change, you can leave it out of the equation. Then, since the atmospheric level is 750 mm.*

$$785 \text{ mmHg} - 75 = 710 \text{ mmHg}$$

5. If a gas at a pressure of 150 mm in a 1.0 L container has its pressure increased to 400 mm, what is the new volume of the container? (Assume temperature is constant).

$$P_1 V_1 = P_2 V_2$$

$$(150 \text{ mm} \times 1.0 \text{ L}) = 400 \text{ mm} \times V_2$$

$$V_2 = 0.375 \text{ L} \rightarrow 0.38 \text{ L}$$

6. What is the new pressure if a 5.00 closed container with a pressure of 2.5 atm is increased to 8.0 L?

$$P_1 V_1 = P_2 V_2$$

$$(2.5 \text{ atm} \times 5.00 \text{ L}) = P_2 \times 8.0 \text{ L}$$

$$P_2 = 1.5625 \rightarrow 1.6 \text{ atm}$$

7. When 0.5 L of oxygen and 4.0 L of Nitrogen both at 0°C and 1.0 atm, are pumped into an evacuated 5.0 L rigid container, what is the final pressure in the container at 0°C? *Since temperature doesn't change, you can leave it out of the equation.*

*So, final pressure of Nitrogen:*

$$P_1 V_1 = P_2 V_2$$

$$(1.0 \text{ atm} \times 4.0 \text{ L}) = P_2 \times 5.0 \text{ L}$$

$$P_2 = 0.8 \text{ atm}$$

*So, final pressure of Oxygen:*

$$P_1 V_1 = P_2 V_2$$

$$(1.0 \text{ atm} \times 0.5 \text{ L}) = P_2 \times 5.0 \text{ L}$$

$$P_2 = 0.1 \text{ atm}$$

$$\text{Total Pressure} = P_1 + P_2 + P_3$$

$$1.0 = 0.8 + 0.1 + P_3 \rightarrow 0.1 \text{ atm}$$

*Atmosphere would combine the two values:*

$$P_1 V_1 = P_2 V_2$$

$$(0.8 \text{ atm} \times 4.0 \text{ L}) = P_2 \times 5.0 \text{ L}$$

$$P_2 = 0.64 \text{ atm}$$

*Then, since pressure volume and moles are directly proportional!*

8. A sealed flask contains a gas with pressure P. If the number of moles of the gas is doubled, the absolute temperature is doubled and the volume of the container is quadrupled, what is the new pressure in terms of P?

*Doubling the moles doubles the pressure.*

$$P_1 V_1 = P_2 V_2$$

*Quadrupling the volume cuts the pressure in 1/4.*

$$P_1 V_1 = P_2 V_2$$

*moles do no overall change in total P.*

9. If 4 moles of gas are added to a container that already holds 1 mole of gas, how will the pressure change within the container?

$$P_1 V_1 = P_2 V_2$$

*Going from 1 mole to 5 moles results in an increase of pressure by 4 times.*

$$P_1 V_1 = P_2 V_2$$

10. If a balloon containing 1000 L of gas at 40.°C and 760 mmHg rises to an altitude where the pressure is 300 mmHg and the temperature is 10.0°C, what is the new volume of the balloon (in L)?

$$\frac{P_1 V_1}{T_1} = \frac{P_2 V_2}{T_2} \times \frac{1 \text{ L}}{1 \text{ L}}$$

$$(760 \text{ mmHg} \times 1000 \text{ L}) / 313 \text{ K} = (300 \text{ mmHg} \times V_2) / 273 \text{ K}$$

$$V_2 = 651.76 \rightarrow 650 \text{ L}$$

11. A sample of nitrogen gas has a pressure of 87.5 mmHg in a 500. mL flask. What is the pressure of this gas sample when it is transferred to a 225.0 mL flask at the same temperature?

$$P_1 V_1 = P_2 V_2$$

$$(87.5 \text{ mmHg} \times 500 \text{ mL}) = P_2 \times 225 \text{ mL}$$

$$P_2 = 194.44 \rightarrow 194 \text{ mmHg}$$

12. You have 4.5 L of  $\text{NO}_2$  at a temperature of 23.0°C. What volume would the  $\text{NO}_2$  occupy at 47°C? (Assume the pressure is constant).

$$\frac{V_1}{T_1} = \frac{V_2}{T_2} \times \frac{1 \text{ L}}{1 \text{ L}}$$

$$(4.5 \text{ L} \times 295.15 \text{ K}) = V_2 \times 320 \text{ K}$$

$$V_2 = 4.225 \rightarrow 4.2 \text{ L}$$

13. You have a sample of a gas in a flask with a volume of 240 mL. At 35.0°C the pressure of the gas is 560 mmHg. If you decrease the temperature to -1.0°C, what is the gas pressure at the lower temperature? *Since pressure doesn't change, you can leave it out of the equation.*

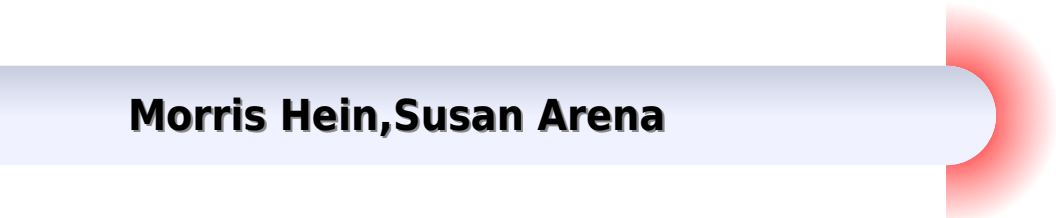
$$P_1 V_1 = P_2 V_2$$

$$(560 \text{ mmHg} \times 240 \text{ mL}) = P_2 \times 240 \text{ mL}$$

$$P_2 = 560 \text{ mmHg} \rightarrow 560 \text{ mmHg}$$

# Unit Gas Laws Homework Chemistry Answers

**Morris Hein, Susan Arena**



## Unit Gas Laws Homework Chemistry Answers:

Introductory Chemistry Problems: The Chemistry And Physics Behind The Numbers Michael E Green, 2025-07-18 The essential point of the book is to show students in Introductory Chemistry classes how to obtain answers that are at least reasonable and to realize when there must have been an error because the answer falls outside the range of possible values. Students often leave ridiculous answers because they simply plug into equations they have memorized correctly or incorrectly. They have not learned to think about the meaning of the numbers that appear in their calculations. This book attempts to show students how to correct this.

*Homework Helpers: Chemistry* Greg Curran, 2025-09-12 Homework Helpers Chemistry is a user friendly review book that will make every student or parent trying to help their child feel like he or she has a private Chemistry tutor. Concepts are explained in clear easy to understand language and problems are worked out with step by step methods that are easy to follow. Each lesson comes with numerous review questions and answer keynotes that explain each correct answer and why it is correct. This book covers all of the topics in a typical one year Chemistry curriculum including A systematic approach to problem solving conversions and the use of units Naming compounds writing formulas and balancing chemical equations Gas laws chemical kinetics acids and bases electrochemistry and more. While Homework Helpers Chemistry is an excellent review for any standardized Chemistry test including the SAT II its real value is in providing support and guidance during the year's entire course of study.

Chemistry John Kenkel, Paul B. Kelter, David S. Hage, 2000-09-21 What a great idea an introductory chemistry text that connects students to the workplace of practicing chemists and chemical technicians. Tying chemistry fundamentals to the reality of industrial life. Chemistry An Industry Based Introduction with CD ROM covers all the basic principles of chemistry including formulas and names chemical bonds.

**Chemistry** Carson-Dellosa Publishing, 2015-03-16 Chemistry for grades 9 to 12 is designed to aid in the review and practice of chemistry topics. Chemistry covers topics such as metrics and measurements matter atomic structure bonds compounds chemical equations molarity and acids and bases. The book includes realistic diagrams and engaging activities to support practice in all areas of chemistry. The 100 Series science books span grades 5 to 12. The activities in each book reinforce essential science skill practice in the areas of life science physical science and earth science. The books include engaging grade appropriate activities and clear thumbnail answer keys. Each book has 128 pages and 100 pages or more of reproducible content to help students review and reinforce essential skills in individual science topics. The series will be aligned to current science standards.

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Basic Chemistry Concepts and Exercises John Kenkel,2010-10-04 Chemistry can be a daunting subject for the uninitiated and all too often introductory textbooks do little to make students feel at ease with the complex subject matter Basic Chemistry Concepts and Exercises brings the wisdom of John Kenkel s more than 35 years of teaching experience to communicate the fundamentals of chemistry in a practical d **Illinois Chemistry Teacher** ,1998 Chemical Principles Richard Earl Dickerson,1984 The Software Encyclopedia 2001 ,2001 *The Educational Value of Chemical Demonstrations in the College Prep Chemistry Classroom* Katherine E. Hagerman,2010 *Advancing the STEM Agenda* Cindy P. Veenstra,Fernando F. Padró,Julie A. Furst-Bowe,2012-05-15 In July 2011 the ASQ Education Division held its first Advancing the STEM Science Technology Engineering and Mathematics Agenda in Education the Workplace and Society Conference at the University of Wisconsin Stout This publication is a selection of papers and workshops from this groundbreaking conference The ideas presented here will help other educators and policy makers to develop their own innovative high impact ideas for inspiring student interest in STEM careers improving the delivery of STEM education at their schools and colleges and helping STEM college graduates transition to the workplace The chapters in this book reflect research and best practices integrating the ideas of continuous improvement in combination with a can do attitude to provide a valuable resource that will lead others to consider similar innovative and collaborative educational structures that will drive more interest in STEM majors in college and provide for our next generation of scientists technicians and engineers Prior to reviewing Advancing the STEM Agenda I had a list in my mind of topics that I hoped would be addressed I m very pleased with how many are covered and covered well This project succeeds at the challenge of providing not only beneficial breadth but also important depth Because our public private partnership has been committed explicitly to continuous improvement for more than a decade I couldn t help but notice as the editors also point out in their conclusion the extent to which

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