

1. Which quantities must be conserved in all chemical reactions?

- A) mass, charge, density
☒ B) mass, charge, energy
C) charge, volume, density
D) charge, volume, energy

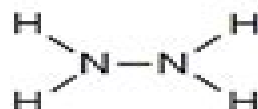
2. Given the balanced equation:



Which statement describes the process represented by this equation?

- A) A bond is formed as energy is absorbed.
☒ B) A bond is formed and energy is released.
C) A bond is broken as energy is absorbed.
D) A bond is broken and energy is released.

3. Given the formula for hydrazine:



How many pairs of electrons are shared between the two nitrogen atoms?

- ☒ A) 1 B) 2 C) 3 D) 4

4. What is the number of pairs of electrons shared in a molecule of N_2 ?

- A) 1 B) 2 ☒ C) 3 D) 6

5. Which atom in the ground state has a stable valence electron configuration?

- ☒ A) Ar B) Al C) Si D) Na

6. An atom of which element has the strongest attraction for electrons in a chemical bond?

- ☒ A) chlorine B) carbon
C) phosphorus D) sulfur

7. Which bond is *least* polar?

- A) As-Cl B) Bi-Cl
☒ C) P-Cl ☒ D) N-Cl

8. Given the electron dot diagram:



The electrons in the bond between hydrogen and fluorine are more strongly attracted to the atom of

- A) hydrogen, which has the higher electronegativity
☒ B) fluorine, which has the higher electronegativity
C) hydrogen, which has the lower electronegativity
D) fluorine, which has the lower electronegativity

9. Given the reaction:



The metal represented by M is most likely a metal from Group

- A) 1 ☒ B) 2 C) 11 D) 13

10. Which type of substance can conduct electricity in the liquid phase but *not* in the solid phase?

- ☒ A) ionic compound
B) molecular compound
C) metallic element
D) nonmetallic element

11. The chemical bonding in sodium phosphate, Na_3PO_4 , is classified as

- A) ionic, only
B) metallic, only
☒ C) both covalent and ionic
D) both covalent and metallic

12. Which type of bonding is found in all molecular substances?

- ☒ A) covalent bonding B) hydrogen bonding
C) ionic bonding D) metallic bonding

13. Which type of substance is soft, has a low melting point, and is a poor conductor of heat and electricity?

- A) network solid ☒ B) molecular solid
C) metallic solid D) ionic solid

Unit 6 Bonding Practice Test Regents Chemistry

Tobias Bleicker



Unit 6 Bonding Practice Test Regents Chemistry:

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