

Limiting and Excess Reactant, Theoretical and Percent Yield.

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8:51 pm

To find the excess reactant and limiting reactant, you can divide the amount of given moles with their coefficient in the balanced equation. The one that has a lower coefficient to mole ratio is the limiting reactant.

Example:



Another way is to compute their theoretical yield, the one that gives the smaller amount of product is the limiting reagent.

2 moles oxygen & 2 mol propane

$$\begin{aligned} & \frac{2 \text{ mol C}_3\text{H}_8}{1 \text{ mol C}_3\text{H}_8} \cdot \frac{3 \text{ mol CO}_2}{1 \text{ mol C}_3\text{H}_8} = 6 \text{ mol} \\ & \frac{2 \text{ mol O}_2}{5 \text{ mol O}_2} \cdot \frac{3 \text{ mol CO}_2}{1 \text{ mol O}_2} = 1.2 \text{ mol} \end{aligned}$$

How to find the percent yield.

$$\% \text{ yield} = \frac{\text{actual yield}}{\text{theoretical yield}} \times 100\%$$

If instead of 4.8 mol of carbon dioxide you got 4.5, what would be the percent yield?

$$\% \text{ yield} = \frac{4.5 \text{ mol}}{4.8 \text{ mol}} \times 100\% = 93.75\%$$

Then compute the error yield by subtracting the percent yield by 100.

$$100 - 93.75 = 6.25\%$$

How to convert ppm to molarity

$$\text{ppm} \frac{\text{g solute}}{\text{g soln}} \times 10^6 \quad M = \frac{\text{mol of solute}}{\text{L of soln}}$$

example

A solution consisting of water and NaOH has a density of 1.05 g/mL and a solute concentration of 840 ppm. What is the molarity of this solution?

To convert grams of solution to L of solution we need density. As well as needing the Molar Mass to convert grams to moles.

$$\text{NaOH} = 39.998 = 40 \text{ g/mol}$$

Step by step process:

$$\text{ppm} = \frac{\text{g solute}}{\text{g solution}} \cdot 1 \times 10^{-6} \quad M = \frac{\text{mol of solute}}{\text{L of solution}}$$

1. First we will write down the ppm, to convert to the grams to moles.

$$\frac{840 \text{ g NaOH}}{1 \times 10^6 \text{ g soln}} \cdot \frac{1 \text{ mol}}{40 \text{ g NaOH}} = \frac{21 \text{ mol}}{1 \times 10^6 \text{ g soln}}$$

2. Then, we will multiply what we got to the substance's density to convert g/1x10⁶g to g/mL.

$$\frac{21 \text{ mol}}{1 \times 10^6 \text{ g}} \cdot \frac{1.05 \text{ g}}{\text{mol}} = 22$$

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Lifelique



Unit 4 Limiting Reactant And Percent Yield Ws 7:

Theoretical Chemical Engineering Abstracts ,1980 **Limiting Reactant and Percent** Lifeliqe,2019 This lesson plan covers analyzing chemical reactions in order to determine limiting reactants and excess reactants including calculating the amount of excess reactant calculate the theoretical yield of a reaction and calculating the percent yield of a reaction

Unit 4 Limiting Reactant And Percent Yield Ws 7 Book Review: Unveiling the Power of Words

In some sort of driven by information and connectivity, the ability of words has are more evident than ever. They have the ability to inspire, provoke, and ignite change. Such may be the essence of the book **Unit 4 Limiting Reactant And Percent Yield Ws 7**, a literary masterpiece that delves deep in to the significance of words and their affect our lives. Written by a renowned author, this captivating work takes readers on a transformative journey, unraveling the secrets and potential behind every word. In this review, we shall explore the book is key themes, examine its writing style, and analyze its overall impact on readers.

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