

Assessment #3 Review

Name: _____

You must show all work for credit.

Due Date: _____ (No late packets will be accepted)

I. Balance Equations*Balance the equations*

- $\text{Mg} + \underline{2} \text{H}_2\text{O} \rightarrow \text{Mg}(\text{OH})_2 + \text{H}_2$
- $\text{Mg} + \underline{2} \text{NaBr} \rightarrow \text{MgBr}_2 + \underline{2} \text{Na}$
- $\text{C}_3\text{H}_8 + \underline{5} \text{O}_2 \rightarrow \underline{3} \text{CO}_2 + \underline{4} \text{H}_2\text{O}$
- $\underline{4} \text{P} + \underline{5} \text{O}_2 \rightarrow \underline{2} \text{P}_2\text{O}_5$
- $\text{Fe}_2(\text{SO}_4)_3 + \underline{6} \text{KOH} \rightarrow \underline{3} \text{K}_2\text{SO}_4 + \underline{2} \text{Fe}(\text{OH})_3$
- $\text{Fe}_2\text{O}_3 + \underline{3} \text{H}_2 \rightarrow \text{Fe} + \underline{3} \text{H}_2\text{O}$

7. How many moles of hydrogen gas react with 1 mole of Fe_2O_3 in reaction #6? 3

Write the balanced equation for problems # 8 and 9:

8. Hydrogen gas and nitrogen monoxide combine to form water and nitrogen gas



9. When sodium phosphate and calcium chloride are mixed they form calcium phosphate and sodium chloride



10. Using the following balanced equation, prove (or explain) the Law of Conservation of Mass.



2 N on left, 2 N on right, 6 H on left, 6 H on right, 2 on left, 2 on right.
sub created, or lost

11. Using the balanced equation shown in problem #10, explain what the coefficients represent.

2 N, 6 H, 2 on left, 2 on right.
sub created, or lost

12. Using the balanced equation shown in problem #10, list the products of the reaction.

2 NH₃

13. Using the balanced equation shown in problem #10, list the reactants of the reaction.

N₂, 3 H₂

II. Types of Reactions*Directions: Identify the type of reaction.*

S = synthesis

SD = single displacement

C = combustion

D = decomposition

DD = double displacement



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Unit 6 Test Chemistry:

Unveiling the Energy of Verbal Beauty: An Emotional Sojourn through **Unit 6 Test Chemistry**

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Table of Contents Unit 6 Test Chemistry

1. Understanding the eBook Unit 6 Test Chemistry
 - The Rise of Digital Reading Unit 6 Test Chemistry
 - Advantages of eBooks Over Traditional Books
2. Identifying Unit 6 Test Chemistry
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Unit 6 Test Chemistry
 - User-Friendly Interface
4. Exploring eBook Recommendations from Unit 6 Test Chemistry
 - Personalized Recommendations

- Unit 6 Test Chemistry User Reviews and Ratings
- Unit 6 Test Chemistry and Bestseller Lists
- 5. Accessing Unit 6 Test Chemistry Free and Paid eBooks
 - Unit 6 Test Chemistry Public Domain eBooks
 - Unit 6 Test Chemistry eBook Subscription Services
 - Unit 6 Test Chemistry Budget-Friendly Options
- 6. Navigating Unit 6 Test Chemistry eBook Formats
 - ePub, PDF, MOBI, and More
 - Unit 6 Test Chemistry Compatibility with Devices
 - Unit 6 Test Chemistry Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Unit 6 Test Chemistry
 - Highlighting and Note-Taking Unit 6 Test Chemistry
 - Interactive Elements Unit 6 Test Chemistry
- 8. Staying Engaged with Unit 6 Test Chemistry
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Unit 6 Test Chemistry
- 9. Balancing eBooks and Physical Books Unit 6 Test Chemistry
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Unit 6 Test Chemistry
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Unit 6 Test Chemistry
 - Setting Reading Goals Unit 6 Test Chemistry
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Unit 6 Test Chemistry
 - Fact-Checking eBook Content of Unit 6 Test Chemistry

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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