

Chemistry – Unit 5 Worksheet 2

1. An old (pre-1987) penny is nearly pure copper. If such a penny has a mass of 3.3 g, how many moles of copper atoms would be in one penny?

$$3.3 \text{ g} \times \frac{1 \text{ mole}}{63.5 \text{ g}} = 5.2 \times 10^{-2} \text{ mol or } 0.052 \text{ mol}$$

2. Four nails have a total mass of 4.42 grams. How many moles of iron atoms do they contain?

$$4.42 \text{ g} \times \frac{1 \text{ mole}}{55.8 \text{ g}} = 7.92 \times 10^{-2} \text{ mol or } 0.0792 \text{ mol}$$

3. A raindrop has a mass of 0.050 g. How many moles of water does a raindrop contain?

$$0.050 \text{ g} \times \frac{1 \text{ mole}}{18.0 \text{ g}} = 2.78 \times 10^{-3} \text{ mol} \rightarrow 2.8 \times 10^{-3} \text{ mol}$$

4. What mass of water would you need to have 15.0 moles of H₂O?

$$15.0 \text{ moles} \times \frac{18.0 \text{ g}}{1 \text{ mole}} = 2.70 \times 10^2 \text{ g}$$

5. One box of Morton's Salt contains 737 grams. How many moles of sodium chloride is this?

$$737 \text{ g} \times \frac{1 \text{ mole}}{58.5 \text{ g}} = 12.6 \text{ mol}$$

Na	23.0
Cl	35.5
	58.5 g

6. A chocolate chip cookie recipe calls for 0.050 moles of baking soda (sodium bicarbonate). How many grams should the chef mass out?

$$0.050 \text{ moles} \times \frac{84.0 \text{ g}}{1 \text{ mole}} = 4.2 \text{ g}$$

Na	23.0
H	1.0
C	12.0
O	48.0
	84.0

7. Rust is iron(III) oxide. The owner of a 1959 Cadillac convertible wants to restore it by removing the rust with oxalic acid, but he needs to know how many moles of rust will be involved in the reaction. How many moles of iron(III) oxide are contained in 2.50 kg of rust?

$$2.50 \text{ kg} \times \frac{1000 \text{ g}}{1 \text{ kg}} \times \frac{1 \text{ mole}}{159.6 \text{ g}} = 15.7 \text{ mol Fe}_2\text{O}_3$$

Fe	111.6
O	48.0
	159.6

Unit 9 2 Modeling Chemistry Answers

SB Merriam



Unit 9 2 Modeling Chemistry Answers:

Chemistry in the Community American Chemical Society, 2006-01-31 This laboratory based text centres itself around decision making activities where students apply their chemistry knowledge to realistic situations This fifth edition includes more photographs new drawings and new design Pharmaceutical Chemistry [GPAT] – Books [Study Notes] 3 in 1 Books with 2000+ Question Answer As Per Updated Syllabus DIWAKAR EDUCATION HUB, 2022-04-01 Pharmaceutical Chemistry GPAT Books Study Notes 3 Books with 2000 Question Answer As Per Updated Syllabus Design by Expert Faculties for Secure 152 Marks in Graduate Pharmacy Aptitude Test Asked 38 MCQ in Exam Highlights of Books As Per Updated Syllabus Graduate Pharmacy Aptitude Test 3 Booklets theory MCQ In Each Book given 6 to 7 Chapters in Details Total 14 Covered Two Types of Chemistry 1 Pharmaceutical Inorganic Chemistry 2 Medicinal Chemistry Total 2000 Questions Answer Numerical with Explanation Design by Pharma Professor Topper Qualified Students Total 3 Booklets For Secured 152 Marks in Exam For More Details Call Whats App 7310762592 7078549303 **Cumulated Index Medicus**, 1999 Index Medicus, 2004 **Developing Solid Oral Dosage Forms** Yihong Qiu, Yisheng Chen, Geoff G.Z. Zhang, Lawrence Yu, Rao V. Mantri, 2016-11-08 Developing Solid Oral Dosage Forms Pharmaceutical Theory and Practice Second Edition illustrates how to develop high quality safe and effective pharmaceutical products by discussing the latest techniques tools and scientific advances in preformulation investigation formulation process design characterization scale up and production operations This book covers the essential principles of physical pharmacy biopharmaceutics and industrial pharmacy and their application to the research and development process of oral dosage forms Chapters have been added combined deleted and completely revised as necessary to produce a comprehensive well organized valuable reference for industry professionals and academics engaged in all aspects of the development process New and important topics include spray drying amorphous solid dispersion using hot melt extrusion modeling and simulation bioequivalence of complex modified released dosage forms biowaivers and much more Written and edited by an international team of leading experts with experience and knowledge across industry academia and regulatory settings Includes new chapters covering the pharmaceutical applications of surface phenomenon predictive biopharmaceutics and pharmacokinetics the development of formulations for drug discovery support and much more Presents new case studies throughout and a section completely devoted to regulatory aspects including global product regulation and international perspectives **Resources in Education**, 1998-10 **Introduction to Geochemistry** Kula C. Misra, 2012-05-21 This book is intended to serve as a text for an introductory course in geochemistry for undergraduate graduate students with at least an elementary level background in earth sciences chemistry and mathematics The text containing 83 tables and 181 figures covers a wide variety of topics ranging from atomic structure to chemical and isotopic equilibria to modern biogeochemical cycles which are divided into four interrelated parts Crystal Chemistry Chemical Reactions and biochemical reactions involving bacteria Isotope Geochemistry radiogenic and stable

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EPA Publications Bibliography United States. Environmental Protection Agency, 1995

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Hydrometallurgy Michael Free, 2013-10-07 This book provides a college level overview of chemical processing of metals in water based solutions in the field that is known as hydrometallurgy

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