

Stoichiometry "End of Chapter" IN CLASS Practice Assessment

Background Info -

- Why are the coefficients in a balanced equation necessary for stoichiometry?
Coefficients are necessary for establishing a molar/mole ratio to convert between moles of different substances.

Skill - Using a balanced equation, calculate using mole to mole ratios. Using a balanced equation, given the grams of reactant, call. the grams of product

- Consider the following unbalanced equation.



There are 5.56g of manganese(IV) oxide present, calculate the grams of chlorine gas.

$$5.00\text{g MnO}_2 \times \frac{1\text{mol MnO}_2}{86.94\text{g MnO}_2} \times \frac{1\text{mol Cl}_2}{1\text{mol MnO}_2} \times \frac{70.90\text{g Cl}_2}{1\text{mol Cl}_2} = 4.00\text{g Cl}_2$$

Skill - After calculating theoretical yield in grams, calculate the percent yield

- You are assigned to complete a lab. Before the reaction is complete, you calculate that a reaction should produce 4.0g Cu(s). After the lab is over you find that 2.8g Cu(s) was collected. What would the percent yield be for that reaction completed within the lab?

$$\% \text{ yield} = \left[\frac{\text{actual/theoretical}}{1} \right] \times 100 \\ = \left[\frac{2.8\text{g}/4.0\text{g}}{1} \right] \times 100 = 70\%$$

Skill - Given experimental data, determine the limiting reactant and the excess reactant

- Calcium carbonate combines with gaseous hydrogen chloride to produce solid calcium chloride, gaseous carbon dioxide, and water vapor. What is the percent yield if 155g of calcium carbonate combines with 250.0g of hydrogen chloride and only 142.0g of calcium chloride is actually obtained within the lab setting? Complete and show work for all 3 steps of a stoichiometry problem.



$$\begin{array}{l} 155\text{g} \\ 155\text{g CaCO}_3 \times \frac{1\text{mol CaCO}_3}{100.09\text{g CaCO}_3} = 1.55\text{mol CaCO}_3 \quad /1\text{mol in rxn} = 1.55 \text{ mols} \\ 250\text{g HCl} \times \frac{1\text{mol HCl}}{36.45\text{g HCl}} = 6.857 \text{ mol HCl} \quad /2\text{mol in rxn} = 3.43 \text{ mols} \end{array}$$

$$1.55\text{mol CaCO}_3 \times \left(\frac{1\text{mol CaCl}_2}{1\text{mol CaCO}_3} \right) \times \left(\frac{110.98\text{g CaCl}_2}{1\text{mol CaCl}_2} \right) = 172.02\text{g}$$

$$\% \text{ yield} = \left[\frac{\text{actual/theoretical}}{1} \right] \times 100 \\ = \left[\frac{142.0\text{g}/172.02\text{g}}{1} \right] \times 100 = 82.55\%$$

Unit 9 Stoichiometry Test Answer

Clemens Wendtner



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Solving Problems in Chemistry Rod O'Connor, Charles Mickey, Alton Hassell, 1977 **A Stoichiometry Unit** David Callaghan, 2004 *Neet Chapter-Wise & Topic-Wise Solved Papers: Chemistry (2005-2022) With 5 Mock Test* Dr. K. G. Ojha, 2022-06-02 The knowledge of Chemistry helps you to understand the world around you From food to pharmaceuticals Chemistry plays a huge role in making informed decisions Therefore to brush up your intellect we present the NEET Chapterwise and Topicwise Chemistry Solved Papers 2005 2021 which is designed to provide a simplified yet systematic understanding to ace the examination The Study Material is strictly based on NCERT Latest Exam Solved Paper is included The Concepts are explained in depth Chapters are compiled with Previous Years Questions Answers to Questions included with Explanations Presence of accurate Figures throughout Five sets of Mock Tests are also included at the end This title focuses on an all inclusive preparations providing the aspirants to learn revise test and gauge their progress against the examination level The Book contains the following units Unit I Physical Chemistry I Unit II Physical Chemistry II Unit III Organic Chemistry I Unit IV Organic Chemistry II Unit V Inorganic Chemistry I Unit VI Inorganic Chemistry II

Stoichiometry Unit Project Luann Marie Decker, 1998 AP Chemistry For Dummies Peter J. Mikulecky, Michelle Rose Gilman, Kate Brutlag, 2008-11-13 A practical and hands on guide for learning the practical science of AP chemistry and preparing for the AP chem exam Gearing up for the AP Chemistry exam AP Chemistry For Dummies is packed with all the resources and help you need to do your very best Focused on the chemistry concepts and problems the College Board wants you to know this AP Chemistry study guide gives you winning test taking tips multiple choice strategies and topic guidelines as well as great advice on optimizing your study time and hitting the top of your game on test day This user friendly guide helps you prepare without perspiration by developing a pre test plan organizing your study time and getting the most out of your AP course You ll get help understanding atomic structure and bonding grasping atomic geometry understanding how colliding particles produce states and so much more To provide students with hands on experience AP chemistry courses include extensive labwork as part of the standard curriculum This is why the book dedicates a chapter to providing a brief review of common laboratory equipment and techniques and another to a complete survey of recommended AP chemistry experiments Two full length practice exams help you build your confidence get comfortable with test formats identify your strengths and weaknesses and focus your studies You ll discover how to Create and follow a pretest plan Understand everything you must know about the exam Develop a multiple choice strategy Figure out displacement combustion and acid base reactions Get familiar with stoichiometry Describe patterns and predict properties Get a handle on organic chemistry nomenclature Know your way around laboratory concepts tasks equipment and safety Analyze laboratory data Use practice exams to maximize your score Additionally you ll have a chance to brush up on the math skills that will help you on the exam learn the critical types of chemistry problems and become familiar with the annoying exceptions to chemistry rules Get your

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