

**Classifying and Balancing Equations**  
Multiple Choice

Name \_\_\_\_\_  
Date \_\_\_\_\_

Per \_\_\_\_\_

**PRACTICE TEST**

- \_\_\_\_\_ 1. A chemical reaction is a process in which  
a. products change into reactants.      c. substances can change state  
b. the law of conservation of mass applies      d. all of these
- \_\_\_\_\_ 2. During a chemical reaction,  
a. new elements are produced      c. atoms are destroyed  
b. atoms are rearranged      d. elements are destroyed
- \_\_\_\_\_ 3. An equation is balanced by  
a. changing subscripts      c. erasing elements as necessary  
b. adding coefficients      d. adding elements as necessary
- \_\_\_\_\_ 4. An atom's ability to undergo chemical reactions is determined by its  
a. protons      b. innermost electrons      c. neutrons      d. valence electrons
- \_\_\_\_\_ 5. What are the reactants in the following chemical equation:  
$$\text{Zn} + \text{CuSO}_4 \longrightarrow \text{ZnSO}_4 + \text{Cu}$$
  
a. zinc and copper      c. zinc and copper (II) sulfate  
b. zinc sulfate and copper      d. only zinc
- \_\_\_\_\_ 6. What are the products in the above equation?  
a. zinc and copper      c. zinc and copper (II) sulfate  
b. zinc sulfate and copper      d. only zinc

Short Answer

For questions 7-12, classify the reaction according to the type it is. Put that answer in the blank. Then add coefficients to balance the reaction when necessary.

- \_\_\_\_\_ 7.  $\text{Zn} + \text{H}_2\text{SO}_4 \longrightarrow \text{ZnSO}_4 + \text{H}_2$
- \_\_\_\_\_ 8.  $\text{H}_2\text{CO}_3 \longrightarrow \text{CO}_2 + \text{H}_2\text{O}$
- \_\_\_\_\_ 9.  $\text{CaCO}_3 + \text{HCl} \longrightarrow \text{CaCl}_2 + \text{H}_2\text{CO}_3$
- \_\_\_\_\_ 10.  $\text{AgNO}_3 + \text{Zn} \longrightarrow \text{Zn}(\text{NO}_3)_2 + \text{Ag}$
- \_\_\_\_\_ 11.  $\text{C}_3\text{H}_8 + \text{O}_2 \longrightarrow \text{CO}_2 + \text{H}_2\text{O}$
- \_\_\_\_\_ 12.  $\text{C}_2\text{H}_5\text{OH} + \text{O}_2 \longrightarrow \text{CO}_2 + \text{H}_2\text{O}$

Write a balanced equation for each of the following reactions:

13. Magnesium chloride is the product of a reaction between magnesium and chlorine.

14. Copper (II) hydroxide and potassium sulfate are produced when potassium hydroxide reacts with copper (II) sulfate.

**Unit 6 Practice Test – Key**

1. b  
2. b  
3. b  
4. d  
5. c  
6. B  
7. Single replacement 1,1,1,1  
8. Decomposition 1,1,1  
9. Double replacement 1,2,1,1-  
   $\text{CaCO}_3$  is ppt (see Ref. Table)  
10. Single replacement 2,1,1,2

11. Combustion 1, 5, 3, 4  
12. Combustion 1,3,2,3  
13.  $\text{Mg} + \text{Cl}_2 \rightarrow \text{MgCl}_2$  1,1,1  
14.  $\text{KOH} + \text{CuSO}_4 \rightarrow \text{Cu}(\text{OH})_2 + \text{K}_2\text{SO}_4$  2,1,1,1

# Unit Test Review Chemical Reactions Answer Key

**DJ Losen**



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