

**AP Chemistry Unit 4: Chemical Reactions**  
**Worksheet 1**

- Calculate the mass percent of oxygen in each of the following compounds:  
(a)  $\text{CuCO}_3$  (b)  $\text{Fe}(\text{HCO}_3)_2$
- Calculate the mass of oxygen in each of the following samples:  
(a) 2.45 g of  $\text{C}_4\text{H}_8\text{O}_3$  (b) 7.10 g of  $\text{Fe}(\text{HCO}_3)_2$
- A certain compound has a formula of  $\text{Cu}_x\text{O}_3$ , where  $x$  is the subscript on the copper ion. The mass percent of the oxygen in a 2.0 g sample is found to be 11.1%.  
(a) How many moles of oxygen are present in the sample?  
(b) What is the molar mass of the compound?  
(c) What is the subscript on the Cu?  
$$6 \text{ Li}(s) + \text{N}_2(g) \rightarrow 2 \text{ Li}_3\text{N}(s)$$
- If 1.50 g of lithium reacts fully via above the reaction:  
(a) How many grams of  $\text{N}_2$  also react?  
(b) How many grams of  $\text{Li}_3\text{N}$  are created?  
$$\text{MgCO}_3(s) + 2 \text{ HCl}(aq) \rightarrow \text{MgCl}_2(aq) + \text{CO}_2(g) + \text{H}_2\text{O}(l)$$
- When the above reaction goes to completion, 245 mL of  $\text{CO}_2$  is produced at STP.  
(a) What mass of  $\text{CO}_2$  is present?  
(b) What mass of  $\text{MgCO}_3$  reacted?  
(c) How many water molecules were produced?  
$$2 \text{ NaN}_3(s) \rightarrow 6 \text{ Na}(s) + \text{N}_2(g)$$
- $3.78 \times 10^{22}$  molecules of  $\text{N}_2$  are created when some  $\text{NaN}_3$  decomposes.  
(a) How many moles of  $\text{NaN}_3$  decomposed?  
(b) How many moles of  $\text{Na}$  are created?

## Unit 4 Section 1 Quiz A Chemistry

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Moore, Peter J. Mikulecky**



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### **Table of Contents Unit 4 Section 1 Quiz A Chemistry**

1. Understanding the eBook Unit 4 Section 1 Quiz A Chemistry
  - The Rise of Digital Reading Unit 4 Section 1 Quiz A Chemistry
  - Advantages of eBooks Over Traditional Books
2. Identifying Unit 4 Section 1 Quiz A Chemistry
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Unit 4 Section 1 Quiz A Chemistry
  - User-Friendly Interface
4. Exploring eBook Recommendations from Unit 4 Section 1 Quiz A Chemistry
  - Personalized Recommendations
  - Unit 4 Section 1 Quiz A Chemistry User Reviews and Ratings
  - Unit 4 Section 1 Quiz A Chemistry and Bestseller Lists

5. Accessing Unit 4 Section 1 Quiz A Chemistry Free and Paid eBooks
  - Unit 4 Section 1 Quiz A Chemistry Public Domain eBooks
  - Unit 4 Section 1 Quiz A Chemistry eBook Subscription Services
  - Unit 4 Section 1 Quiz A Chemistry Budget-Friendly Options
6. Navigating Unit 4 Section 1 Quiz A Chemistry eBook Formats
  - ePub, PDF, MOBI, and More
  - Unit 4 Section 1 Quiz A Chemistry Compatibility with Devices
  - Unit 4 Section 1 Quiz A Chemistry Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Unit 4 Section 1 Quiz A Chemistry
  - Highlighting and Note-Taking Unit 4 Section 1 Quiz A Chemistry
  - Interactive Elements Unit 4 Section 1 Quiz A Chemistry
8. Staying Engaged with Unit 4 Section 1 Quiz A Chemistry
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Unit 4 Section 1 Quiz A Chemistry
9. Balancing eBooks and Physical Books Unit 4 Section 1 Quiz A Chemistry
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Unit 4 Section 1 Quiz A Chemistry
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Unit 4 Section 1 Quiz A Chemistry
  - Setting Reading Goals Unit 4 Section 1 Quiz A Chemistry
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Unit 4 Section 1 Quiz A Chemistry
  - Fact-Checking eBook Content of Unit 4 Section 1 Quiz A Chemistry
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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