Name:

#### Chemistry Practice: Balancing Chemical Equations

 $OO_2$ 

- Balance the equations below. Remember this means only using coeeficients. All formulas are written correctly.
  - a.) 2 NaHCO<sub>3</sub> Na<sub>2</sub>CO<sub>3</sub> + H<sub>2</sub>O +
  - b.) Zn + S ZnS
  - c.) 2H<sub>2</sub> + O<sub>2</sub> → ZH<sub>2</sub>O
  - d.) 2 FeO + C ZFe + CO<sub>2</sub>
  - e.)  $Cl_2 + ZKI \longrightarrow ZKCI + I_2$
  - f.)  $B_2O_3 + 3Mg 3MgO + ZB$

  - (h) 5H<sub>2</sub>O + Na<sub>2</sub>B<sub>4</sub>O<sub>7</sub> + H<sub>2</sub>SO<sub>4</sub> → 4/H<sub>3</sub>BO<sub>3</sub> + Na<sub>2</sub>SO<sub>4</sub>
  - i.)  $CaC_2 + 2H_2O$   $\longrightarrow$   $Ca(OH)_2 + C_2H_2$
  - j.) NH<sub>4</sub>NO<sub>3</sub> N<sub>2</sub>O +2H<sub>2</sub>O
  - k.)  $H_2S + O_2$   $SO_2 + H_2O$
  - H<sub>2</sub>SO<sub>4</sub> + 2NaCl → ZHCl + Na<sub>2</sub>SO<sub>4</sub>

  - n.) 3 BaO + Z Al \_\_\_\_\_ 3 Ba + Al<sub>2</sub>O<sub>3</sub>
  - o.) 3 O<sub>2</sub> -> 2 O<sub>3</sub>
  - p.)  $SiO_2 + ZC$   $\longrightarrow$  ZCO + Si
  - q.) 2 PbS + 3 O<sub>2</sub> \_\_\_\_ 2 PbO + 2 SO<sub>2</sub>
  - r.) 2 HgO → 2 Hg + O<sub>2</sub>
  - s.)  $C_{12}H_{22}O_{11} + /ZO_2 /Z_CO_2 + //H_2O$
  - 1.)  $C_3H_8 + 5O_2 \rightarrow 3CO_2 + 4H_2O$

  - v.) P<sub>2</sub>O<sub>5</sub> + 6 NaOH 2 Na<sub>3</sub>PO<sub>4</sub> + 3 H<sub>2</sub>O
  - w.)  $3 \text{ BaCl}_2 + \text{Al}_2(SO_4)_3 \longrightarrow 3 \text{ BaSO}_4 + 2 \text{AlCl}_3$
  - x.) 4/NH<sub>3</sub> + 5 O<sub>2</sub> 6 H<sub>2</sub>O + 4/NO
  - y.) Z Na<sub>2</sub>O<sub>2</sub> + 2 H<sub>2</sub>O → 4/ NaOH + 4/ O<sub>2</sub>
  - z.)  $Sn(ClO_3)_4$   $\longrightarrow$   $SnCl_4 + 6 O_2$

# **Unit 6 Balancing Chemical Reactions 1 Key**

Randy L. Bell, Joe Garofalo

#### **Unit 6 Balancing Chemical Reactions 1 Key:**

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#### **Table of Contents Unit 6 Balancing Chemical Reactions 1 Key**

- 1. Understanding the eBook Unit 6 Balancing Chemical Reactions 1 Key
  - The Rise of Digital Reading Unit 6 Balancing Chemical Reactions 1 Key
  - Advantages of eBooks Over Traditional Books
- 2. Identifying Unit 6 Balancing Chemical Reactions 1 Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Unit 6 Balancing Chemical Reactions 1 Key
  - User-Friendly Interface
- 4. Exploring eBook Recommendations from Unit 6 Balancing Chemical Reactions 1 Key
  - Personalized Recommendations
  - Unit 6 Balancing Chemical Reactions 1 Key User Reviews and Ratings
  - Unit 6 Balancing Chemical Reactions 1 Key and Bestseller Lists
- 5. Accessing Unit 6 Balancing Chemical Reactions 1 Key Free and Paid eBooks
  - Unit 6 Balancing Chemical Reactions 1 Key Public Domain eBooks
  - Unit 6 Balancing Chemical Reactions 1 Key eBook Subscription Services
  - Unit 6 Balancing Chemical Reactions 1 Key Budget-Friendly Options

- 6. Navigating Unit 6 Balancing Chemical Reactions 1 Key eBook Formats
  - o ePub, PDF, MOBI, and More
  - Unit 6 Balancing Chemical Reactions 1 Key Compatibility with Devices
  - Unit 6 Balancing Chemical Reactions 1 Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - o Adjustable Fonts and Text Sizes of Unit 6 Balancing Chemical Reactions 1 Key
  - Highlighting and Note-Taking Unit 6 Balancing Chemical Reactions 1 Key
  - Interactive Elements Unit 6 Balancing Chemical Reactions 1 Key
- 8. Staying Engaged with Unit 6 Balancing Chemical Reactions 1 Key
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Unit 6 Balancing Chemical Reactions 1 Key
- 9. Balancing eBooks and Physical Books Unit 6 Balancing Chemical Reactions 1 Key
  - $\circ\,$  Benefits of a Digital Library
  - Creating a Diverse Reading Collection Unit 6 Balancing Chemical Reactions 1 Key
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Unit 6 Balancing Chemical Reactions 1 Key
  - Setting Reading Goals Unit 6 Balancing Chemical Reactions 1 Key
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Unit 6 Balancing Chemical Reactions 1 Key
  - Fact-Checking eBook Content of Unit 6 Balancing Chemical Reactions 1 Key
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements

• Interactive and Gamified eBooks

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