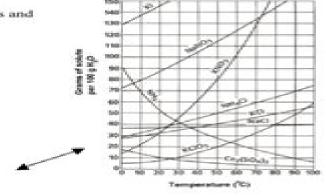
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Unit 12 Solutions Review

- 1. Review learning targets for Unit 12
- 2. Solute, solvent, soln
- 3. Like dissolves like
- 4. Colligative properties (freezing point depression, boiling point elevation, Van't Hoff Factor)
- 5. Saturated, unsaturated, supersaturated
- 6. 4 types of mixtures and examples and examples of each
- What are the differences between the solubilities of gases and solids
- 8. Naming and formula writing
- 9. Steps to make a solution
- 10. Steps to dilute a solution

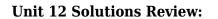
Solution	Supersaturated, Saturated or Unsaturated?	If unsaturated: How much more solute can dissolve in the solution?
a solution that contains 95g of NaNO, at 30°C (in 100 mL H/O)		
a solution that contains 50g of NH ₂ Cl at 50°C (in 100 mL H ₂ O)		
a solution that contains 20g of KCIO; at 50°C (in 100 mL H ₂ O)		
a solution that contains 70g of KI at 0°C (in 100 mL H.Q)		



- 11. Calculate the molarity (M) of a solution prepared by dissolving 175 g of KNO1 in 750 mL of water.
- Calculate the molarity (M) of a solution prepared by dissolving 235.0 g of CaCl₂ in 650.0 mL of water.
- 13. How many mL of 3.0M HCl are needed to make 300.0 mL of 0.10M HCl?

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Donald E. Kieso, Jerry J. Weygandt, Terry D. Warfield



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