

Chemistry – Unit 5 Worksheet 2

1. An old (pre-1987) penny is nearly pure copper. If such a penny has a mass of 3.3 g, how many moles of copper atoms would be in one penny?

$$3.3 \text{ g} \times \frac{1 \text{ mole}}{63.5 \text{ g}} = 5.2 \times 10^{-2} \text{ mol or } 0.052 \text{ mol}$$

2. Four nails have a total mass of 4.42 grams. How many moles of iron atoms do they contain?

$$4.42 \text{ g} \times \frac{1 \text{ mole}}{55.8 \text{ g}} = 7.92 \times 10^{-2} \text{ mol or } 0.0792 \text{ mol}$$

3. A raindrop has a mass of 0.050 g. How many moles of water does a raindrop contain?

$$0.050 \text{ g} \times \frac{1 \text{ mole}}{18.0 \text{ g}} = 2.78 \times 10^{-3} \text{ mol} \rightarrow 2.8 \times 10^{-3} \text{ mol}$$

4. What mass of water would you need to have 15.0 moles of H₂O?

$$15.0 \text{ moles} \times \frac{18.0 \text{ g}}{1 \text{ mole}} = 2.70 \times 10^2 \text{ g}$$

5. One box of Morton's Salt contains 737 grams. How many moles of sodium chloride is this?

$$737 \text{ g} \times \frac{1 \text{ mole}}{58.5 \text{ g}} = 12.6 \text{ mol}$$

Na	23.0
Cl	35.5
	58.5 g

6. A chocolate chip cookie recipe calls for 0.050 moles of baking soda (sodium bicarbonate). How many grams should the chef mass out?

$$0.050 \text{ moles} \times \frac{84.0 \text{ g}}{1 \text{ mole}} = 4.2 \text{ g}$$

Na	23.0
H	1.0
C	12.0
O	48.0
	84.0

7. Rust is iron(III) oxide. The owner of a 1959 Cadillac convertible wants to restore it by removing the rust with oxalic acid, but he needs to know how many moles of rust will be involved in the reaction. How many moles of iron(III) oxide are contained in 2.50 kg of rust?

$$2.50 \text{ kg} \times \frac{1000 \text{ g}}{1 \text{ kg}} \times \frac{1 \text{ mole}}{159.6 \text{ g}} = 15.7 \text{ mol Fe}_2\text{O}_3$$

Fe	111.6
O	48.0
	159.6

Unit 5 Ws 1 Chemistry

KJ Lindholm-Leary



Unit 5 Ws 1 Chemistry:

Unit 5 Ws 1 Chemistry Book Review: Unveiling the Power of Words

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Table of Contents Unit 5 Ws 1 Chemistry

1. Understanding the eBook Unit 5 Ws 1 Chemistry
 - The Rise of Digital Reading Unit 5 Ws 1 Chemistry
 - Advantages of eBooks Over Traditional Books
2. Identifying Unit 5 Ws 1 Chemistry
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Unit 5 Ws 1 Chemistry
 - User-Friendly Interface
4. Exploring eBook Recommendations from Unit 5 Ws 1 Chemistry
 - Personalized Recommendations
 - Unit 5 Ws 1 Chemistry User Reviews and Ratings
 - Unit 5 Ws 1 Chemistry and Bestseller Lists

5. Accessing Unit 5 Ws 1 Chemistry Free and Paid eBooks
 - Unit 5 Ws 1 Chemistry Public Domain eBooks
 - Unit 5 Ws 1 Chemistry eBook Subscription Services
 - Unit 5 Ws 1 Chemistry Budget-Friendly Options
6. Navigating Unit 5 Ws 1 Chemistry eBook Formats
 - ePub, PDF, MOBI, and More
 - Unit 5 Ws 1 Chemistry Compatibility with Devices
 - Unit 5 Ws 1 Chemistry Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Unit 5 Ws 1 Chemistry
 - Highlighting and Note-Taking Unit 5 Ws 1 Chemistry
 - Interactive Elements Unit 5 Ws 1 Chemistry
8. Staying Engaged with Unit 5 Ws 1 Chemistry
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Unit 5 Ws 1 Chemistry
9. Balancing eBooks and Physical Books Unit 5 Ws 1 Chemistry
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Unit 5 Ws 1 Chemistry
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Unit 5 Ws 1 Chemistry
 - Setting Reading Goals Unit 5 Ws 1 Chemistry
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Unit 5 Ws 1 Chemistry
 - Fact-Checking eBook Content of Unit 5 Ws 1 Chemistry
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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