



## Unit: Chemical Bonding Answer Key "Covalent Bonding" - WS #3

**Directions:** Please answer each question using your notes from today's class. ☺

1. What is the difference between a molecular formula, structural formula and an electron dot formula? Give an example of each.

*A molecular formula gives the chemical make-up of one molecule. It does not show atom-to-atom connectivity. Example:  $C_3H_8$*

*A structural formula uses lines instead of dots to show shared pairs of electrons.*

*Example:  $H-H$*

*An electron dot formula uses dots instead of lines to show shared pairs of electrons.*

*Example:  $H:H$*

2. What is the difference between single, double, and triple covalent bonds? How many shared electrons are in each?

*The difference between the types of bonds is: a single bond is two shared electrons, a double is four shared electrons, and a triple is six shared electrons.*

3. Draw the electron dot and structural formulas for the following molecules.

| Substance | Rough copy (work!) | Electron dot                                                                              | Structural                                              |
|-----------|--------------------|-------------------------------------------------------------------------------------------|---------------------------------------------------------|
| a) $Cl_2$ |                    | <i>Note: The electron dot structures have dots in between each atom instead of lines.</i> | $:\ddot{Cl}-\ddot{Cl}:$                                 |
| b) $H_2S$ |                    |                                                                                           | $\begin{array}{c} :\ddot{S}: \\   \\ H \end{array} - H$ |
| c) $BrI$  |                    |                                                                                           | $:\ddot{I}-\ddot{Br}:$                                  |

# Unit 6 Chemical Bonding Answers

**Xiaolong Qi**



## **Unit 6 Chemical Bonding Answers:**

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## **Table of Contents Unit 6 Chemical Bonding Answers**

1. Understanding the eBook Unit 6 Chemical Bonding Answers
  - The Rise of Digital Reading Unit 6 Chemical Bonding Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Unit 6 Chemical Bonding Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Unit 6 Chemical Bonding Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Unit 6 Chemical Bonding Answers

- Personalized Recommendations
  - Unit 6 Chemical Bonding Answers User Reviews and Ratings
  - Unit 6 Chemical Bonding Answers and Bestseller Lists
5. Accessing Unit 6 Chemical Bonding Answers Free and Paid eBooks
    - Unit 6 Chemical Bonding Answers Public Domain eBooks
    - Unit 6 Chemical Bonding Answers eBook Subscription Services
    - Unit 6 Chemical Bonding Answers Budget-Friendly Options
  6. Navigating Unit 6 Chemical Bonding Answers eBook Formats
    - ePub, PDF, MOBI, and More
    - Unit 6 Chemical Bonding Answers Compatibility with Devices
    - Unit 6 Chemical Bonding Answers Enhanced eBook Features
  7. Enhancing Your Reading Experience
    - Adjustable Fonts and Text Sizes of Unit 6 Chemical Bonding Answers
    - Highlighting and Note-Taking Unit 6 Chemical Bonding Answers
    - Interactive Elements Unit 6 Chemical Bonding Answers
  8. Staying Engaged with Unit 6 Chemical Bonding Answers
    - Joining Online Reading Communities
    - Participating in Virtual Book Clubs
    - Following Authors and Publishers Unit 6 Chemical Bonding Answers
  9. Balancing eBooks and Physical Books Unit 6 Chemical Bonding Answers
    - Benefits of a Digital Library
    - Creating a Diverse Reading Collection Unit 6 Chemical Bonding Answers
  10. Overcoming Reading Challenges
    - Dealing with Digital Eye Strain
    - Minimizing Distractions
    - Managing Screen Time
  11. Cultivating a Reading Routine Unit 6 Chemical Bonding Answers
    - Setting Reading Goals Unit 6 Chemical Bonding Answers
    - Carving Out Dedicated Reading Time
  12. Sourcing Reliable Information of Unit 6 Chemical Bonding Answers

- Fact-Checking eBook Content of Unit 6 Chemical Bonding Answers
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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