

Unit 9 Outline- Thermodynamics, Reaction Rates, & Chemical Equilibrium

Chapters 17 & 18

Essential Skills/ State Standards:

- 9-1. Describe how temperature and heat flow in terms of the motion of molecules (or atoms). (CSS: 7a)
 9-2. Know that chemical processes & phase changes can either release energy (exothermic) or absorb energy (endothermic). (CSS: 7b, 7c)
 9-3. Solve problems involving heat flow and temperature changes, using known values of specific heat ($q = m \cdot c \cdot \Delta T$) & latent heat of phase change ($q = m \cdot \Delta H$). (CSS: 7d)
 9-4. Know the rate of reaction is the decrease in concentration of reactants or the increase in concentration of products with time & depends on such factors as: concentration, temperature, and pressure. (CSS: 8a, 8b)
 9-5. Know the definition and role of activation energy & catalyst in a chemical reaction. (CSS: 8d*, 8c)
 9-6. Know how to use Le Chatelier's principle to predict the effect of changes in concentration, temperature, and pressure (CSS: 9a)
 9-7. Know that equilibrium is established when the forward and reverse reaction rates are equal. (CSS: 9b)

Vocabulary

- | | | | |
|---|---|--|--|
| <ul style="list-style-type: none"> • Condensation • Freezing • Vaporization • Melting • Boiling • Sublimation | <ul style="list-style-type: none"> • Deposition • Heat • Temperature • Thermochemistry • Exothermic • Endothermic | <ul style="list-style-type: none"> • Calorimeter • Joule • Specific heat • $q = m \cdot c \cdot \Delta T$ • $q = m \cdot \Delta H_{\text{vaporization}}$ • $q = m \cdot \Delta H_{\text{fusion}}$ | <ul style="list-style-type: none"> • Heat of reaction • Le Chatelier's principle • Equilibrium • Activation Energy • Catalyst |
|---|---|--|--|

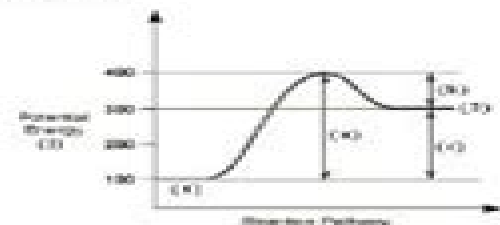
Book Assignments:

Topic	Read	Required Problems
1. Heat, Temperature & Specific Heat	Section 17-1 p. 511-524	Section 17-p. 524 #2, 5 & 6 Chapter 17 Review p. 548 #16 & 17
3. Chemical Equilibrium & Le Chatelier's Principle	Section 12-3 p. 372-375 Section 18-2 p. 562-568	Chapter 18 Review p. 586 #4, 6 & 7
4. Phase Changes Diagrams & Latent Heat of Phase Change	Section 12-3 p. 378-382	Section 12-3 p. 382 #4-8 Chapter 12 Review p. 389 #19, 28, 31 & 37
5. Collision Theory & Activation Energy	Section 17-3 p. 532-534	Section 17-3 p. 537 #2, 3 & 5
6. Reaction Rates & Catalysts	Section 17-4 p. 538-544	Section 17-4 p. 545 #2-4

Study Guide (Use notes, worksheets, and labs in addition to this study guide)

- According to the collision theory, what 3 things must happen in order for two molecules to react?
- What is the importance of the activation energy in a chemical reaction?
- Using the collision theory, explain if the following would increase/ decrease the rate of reaction:

$$2 \text{HCl(aq)} + \text{Mg(s)} \rightarrow \text{MgCl}_2\text{(aq)} + \text{H}_2\text{(g)}$$
 - Increase temp
 - Increase HCl
 - Increase pressure
 - Remove MgCl_2
 - Add H_2
- Identify the following terms with their correct number/letter from the graphs:
 - * Reactants- _____
 - * Products- _____
 - * Activation energy (number/letter & amount in Joules)- _____
 - * Is this reaction exothermic or endothermic?
- Explain how a catalyst affects the activation energy in a chemical reaction. (Also, show on the above graph what a catalyst might do).



Unit 7chemistry Stoichiometry Study Guide Answers

Robert R. Redfield



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Prain,2004-03-31 This volume is of interest to science educators graduate students and classroom teachers The book will also
be an important addition to any scholarly library focusing on science education science literacy and writing This book is
unique in that it synthesizes the research of the three leading researchers in the field of writing to learn science Carolyn S
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which learning occurs It combines both theory and practice in a unique way Although primarily a tool for research classroom
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