

# < KINETIC THEORY >

Introduction: - (Microscopic properties)

(John Dalton)  $\rightarrow$  Atom  $\rightarrow$  1803

In 1873  $\rightarrow$  [Maxwell & Boltzmann]

Ideal Gas  $\rightarrow$  [Kinetic theory of gases]

Atoms/molecules constantly motions  
they are not in rest.

< we study Average properties of gases >

Average Relax time ' $\tau$ '

- Average time period b/w two successive collisions.

< for 1661 >  $\rightarrow$  [Boyle's Law]:-

- Keeping temperature constant, pressure of a given mass of gas varies inversely with volume.  
( $P \propto 1/V$ )

$$\Rightarrow PV = \text{constant}$$

$$\text{or } [P_1 V_1 = P_2 V_2]$$

- At high temperature Real gas and Ideal gas approximately behave same. (High temp. per Boyle law real, Ideal does not follow Boyle law, otherwise diff. Ideal gas.)

< for 1783 >  $\rightarrow$  [Charles's Law]:-

- For a fixed mass pressure, the vol. of a gas is directly proportional to its absolute temperature ' $T$ '.

$$(V \propto T)$$

$$\Rightarrow \frac{V}{T} = \text{constant}$$

$$\text{or } \left[ \frac{V_1}{T_1} = \frac{V_2}{T_2} \right]$$

- At high temperature law holds for all gas (Real + Ideal).

Combining both Boyle's and Charles's law

$$(T = \text{constant}) V \propto \frac{1}{P} \quad \cdot \quad V \propto T \quad (P = \text{constant})$$

$$\Rightarrow V \propto \frac{T}{P} \quad \text{or} \quad \frac{PV}{T} = \text{constant}$$

$$\text{or} \quad \frac{P_1 V_1}{T_1} = \frac{P_2 V_2}{T_2}$$

[Equation of State]

ARIHANT

free path = distance b/w two consecutive collisions

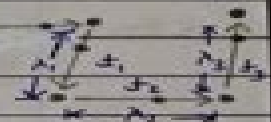
mean free path:-

Average distance travelled by a molecule b/w two successive collisions.

$$\lambda_m = \lambda_1 + \lambda_2 + \dots + \lambda_n$$

(mean)

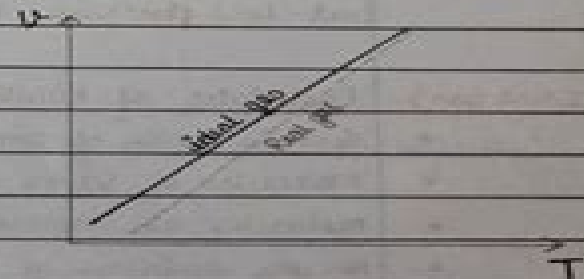
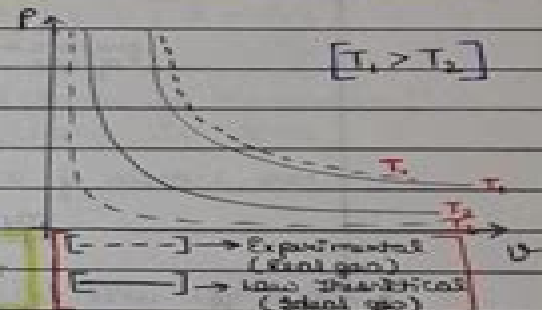
$n$



$$T_m = T_1 + T_2 + \dots + T_n$$

(mean)

$n$



# Study Guide For Kinetic Theory

**Katherine J. Denniston, Denniston  
Katherine**



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