

Name: Date: 

## Student Exploration: Temperature and Particle Motion

**Directions:** Follow the instructions to go through the simulation. Respond to the questions and prompts in the orange boxes.

**Vocabulary:** absolute zero, Kelvin scale, kinetic energy, Maxwell-Boltzmann distribution, molar mass, molecule, temperature, universal gas constant

**Prior Knowledge Questions** (Do these BEFORE using the Gizmo.)

1. Why is hot air hot?

Hot air is hot due to higher pressure and molecules moving faster.

2. Why is cold air cold?

Cold air is cold due to lower pressure and molecules moving slowly.

3. Air consists of tiny particles called **molecules**. How do you think the molecules move in hot and in cold air?

The molecules move fast in hot air and slow in cold air.

### Gizmo Warm-up

The *Temperature and Particle Motion* Gizmo illustrates how the molecules of gas move at different temperatures. In this Gizmo, temperature is measured on the **Kelvin scale**, which measures temperature from **absolute zero**, the coldest possible temperature ( $-273.15^{\circ}\text{C}$ ). Each unit on the Kelvin scale is equivalent to  $1^{\circ}\text{C}$ :  $273.15\text{ K} = 0^{\circ}\text{C}$ , and  $373.15\text{ K} = 100^{\circ}\text{C}$ .

Check that the selected gas is **Hydrogen** and the **Temperature** is 300 K.



1. Describe the motion of the hydrogen molecules:

The molecules move around at a fast pace and are able to roam freely.

2. Are all of the molecules moving at the same speed?

No.

### Activity A:

#### Molecular motions

#### Get the Gizmo ready:

- Check that the selected gas is **Hydrogen** and the **Temperature** is set to 300 K.



**Question:** How is the temperature of a gas related to the motion of gas molecules?

1. **Observe:** Move the **Temperature** slider back and forth. Focus on the particle motion at left. What do you notice?

# Temperature And Particle Motion Gizmo Answer Key

**A Loxley**



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