# Unit 6 Notes: Stoichiometry

Pronounced: Stoy-kee-AHM-uh-tree Name: Key

Stoichiometry: uses mole ratio from BCE to solve problems; branch of chemistry that studies the quantitative relationships in a balanced chemical equation

Stoichiometry is based on the law of conservation of mass. Which states...matter cannot be created or destroyed

How can we interpret coefficients in the following BCE?

4 Al + 3 O2 - 2 Al2O3

number of atoms/comounds 2. number of moles...Mole Ratios!

Mole Ratios: Shows the relationship between moles of reactants and products in a chemical reaction.

Write all possible mole ratios from this equation:

4 Al + 3 O<sub>2</sub> → 2 Al<sub>2</sub>O<sub>3</sub>

3 mol Oa

2 mol Al<sub>2</sub>O<sub>2</sub>

4 mol Al 3 mol O<sub>2</sub> 4 mol Al 3 mol O<sub>3</sub> 4 mol Al 2 mol Al<sub>2</sub>O<sub>3</sub>

2 mol Al<sub>2</sub>O<sub>3</sub> 4 mol Al 2 mol Al<sub>2</sub>O<sub>3</sub> 3 mol O<sub>3</sub>

Mass-Mass Stoichiometry Example

How many grams of oxygen can be produced by decomposing 5.00 g of potassium chlorate?

1) Map it!!

2 KClO<sub>3</sub> → 3 O<sub>2</sub> + 2 KCl

\_\_\_\_5.00g KClO<sub>3</sub> × \_\_1 mol\_\_ × 3 mol O<sub>2</sub> × \_\_32.0 g\_ = 1.96 g O<sub>2</sub> 122.6 g 2 mol KClO<sub>3</sub> 1 mol

Draw an arrow going down from 5g KClO<sub>3</sub> to below KClO<sub>3</sub>, an arrow going to the right to O<sub>2</sub>, arrow going up from below O<sub>2</sub> to the question mark above O<sub>2</sub>

- 2) Get to moles!! (Use the periodic table)
- 3) Use the Mole ratio to get to the new substance.

  (This is stoichiometry, USE THE Balanced Chemical Equation!!)
- 4) Get back into grams. (Again with the periodic table)

# **Unit 10a Stoichiometry Notes**

John A. Ripmeester, Saman Alavi

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