

Unit 7 Chemical Reactions

PRACTICE TEST/ REVIEW

- _____ 1. A chemical reaction is a process in which
- products change into reactants
 - the law of conservation of mass applies
 - substances can change state
 - all of these
- _____ 2. During a chemical reaction,
- new elements are produced
 - atoms are rearranged
 - atoms are destroyed
 - elements are destroyed
- _____ 3. An equation is balanced by
- changing subscripts
 - adding coefficients
 - erasing elements as necessary
 - adding elements as necessary
- _____ 4. What are the reactants in the following chemical equation:
- $$\text{Zn} + \text{CuSO}_4 \rightarrow \text{ZnSO}_4 + \text{Cu}$$
- zinc and copper
 - zinc sulfate and copper
 - zinc and copper (II) sulfate
 - only zinc
- _____ 5. What are the products in the above equation?
- zinc and copper
 - zinc sulfate and copper
 - zinc and copper (II) sulfate
 - only zinc
- _____ 6. The breaking down of a substance into two or simpler substances is _____
- Displacement
 - Synthesis
 - Decomposition
 - Double displacement
- _____ 7. Each substance to the right of the arrow in a chemical equation is a(n)
- Reactant
 - Catalyst
 - Product
 - Inhibitor
- _____ 8. According to the law of conservation of mass, if two atoms of hydrogen are used as a reactant, how many atoms of hydrogen must be part of the product?
- 4
 - 2
 - 1
 - 0
- _____ 9. What type of reaction is shown in the following chemical equation, $\text{NH}_4\text{Cl} \rightarrow \text{NH}_3 + \text{HCl}$?
- Displacement
 - Synthesis
 - Decomposition
 - Double displacement

Unit 7 Chemical Reactions Test

J. G. R. Briggs



Unit 7 Chemical Reactions Test:

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