

Name: _____

Thermochemistry Unit Test

40

Heat capacities:
water: 4.18 J/g°C
ice: 1.96 J/g°C

$$\Delta H_{\text{fus}}(\text{water}) = 6.03 \text{ kJ/mol}$$

True/False (6 marks)

Indicate whether the sentence or statement is true or false.

- _____ 1. According to Hess's Law, if the coefficients of a chemical equation are altered by multiplying or dividing by a constant factor, then the ΔH is *altered the same way but also changes the sign of the ΔH* .
- _____ 2. The following diagram represents an *exothermic reaction*.



- _____ 3. The reactants in a heat of formation reaction can be *elements or compounds* as long as they are in their standard forms.
- _____ 4. Enthalpy changes for an endothermic reaction are given a *negative sign*.
- _____ 5. The amount of energy involved in a change depends on the *quantity* of matter undergoing the change.
- _____ 6. If a chemical equation is reversed, according to Hess's Law, *no change* occurs to the ΔH of the reaction.

Multiple Choice (16 marks)

Identify the letter of the choice that best completes the statement or answers the question.

1. For which one of the following substances is the standard enthalpy of formation, ΔH_f° , equal to zero?
a. water [$\text{H}_2\text{O}_{(l)}$] c. carbon dioxide [$\text{CO}_{2(g)}$] e. tin [$\text{Sn}_{(s)}$]
b. lead [$\text{Pb}_{(s)}$] d. carbon dioxide [$\text{CO}_{2(l)}$]
2. Bond energies for a chemical reaction were used to determine the enthalpy change. If the total energies of breaking the bonds of the reactants was 4678 kJ and that for forming the products was 6870 kJ, the enthalpy change would be: _____
3. The energy required to take 20.0g of ice at -12°C to liquid water at 80°C is: (see given info on p1)
a. 7165 J b. 13.86 kJ c. 7.69 kJ d. 6.69 kJ e. 127.8 kJ
4. Which of the following equations represents the correct heat of formation of water vapour?
a. $\text{H}_{2(g)} + \frac{1}{2} \text{O}_{2(g)} \rightarrow \text{H}_2\text{O}_{(g)}$ b. $\text{H}_{2(g)} + \text{O}_{2(g)} \rightarrow 2 \text{H}_2\text{O}_{(g)}$ c. $\text{H}_{2(g)} + \text{O}_{(g)} \rightarrow \text{H}_2\text{O}_{(g)}$
5. A chemical system in which both energy and matter can flow into or out of a system is described as
a. a closed system b. an open system c. an isolated system
d. a chemical system e. none of the above

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