

Name: _____

Chemistry Practice: Balancing Chemical Equations

1. Balance the equations below. Remember this means only using coefficients. All formulas are written correctly.

- a.) $2 \text{NaHCO}_3 \longrightarrow \text{Na}_2\text{CO}_3 + \text{H}_2\text{O} + \text{CO}_2$
- b.) $\text{Zn} + \text{S} \longrightarrow \text{ZnS}$
- c.) $2 \text{H}_2 + \text{O}_2 \longrightarrow 2 \text{H}_2\text{O}$
- d.) $2 \text{FeO} + \text{C} \longrightarrow 2 \text{Fe} + \text{CO}_2$
- e.) $\text{Cl}_2 + 2 \text{KI} \longrightarrow 2 \text{KCl} + \text{I}_2$
- f.) $\text{B}_2\text{O}_3 + 3 \text{Mg} \longrightarrow 3 \text{MgO} + 2 \text{B}$
- g.) $\text{Ca} + 2 \text{H}_2\text{O} \longrightarrow \text{Ca(OH)}_2 + 2 \text{H}_2$
- h.) $4 \text{H}_2\text{O} + \text{Na}_2\text{B}_4\text{O}_7 + \text{H}_2\text{SO}_4 \longrightarrow 4 \text{H}_3\text{BO}_3 + \text{Na}_2\text{SO}_4$
- i.) $\text{CaC}_2 + 2 \text{H}_2\text{O} \longrightarrow \text{Ca(OH)}_2 + \text{C}_2\text{H}_2$
- j.) $\text{NH}_4\text{NO}_3 \longrightarrow \text{N}_2\text{O} + 2 \text{H}_2\text{O}$
- k.) $\text{H}_2\text{S} + \text{O}_2 \longrightarrow \text{SO}_2 + \text{H}_2\text{O}$
- l.) $\text{H}_2\text{SO}_4 + 2 \text{NaCl} \longrightarrow 2 \text{HCl} + \text{Na}_2\text{SO}_4$
- m.) $2 \text{NaCl} + 2 \text{H}_2\text{O} \longrightarrow 2 \text{NaOH} + \text{Cl}_2 + \text{H}_2$
- n.) $3 \text{BaO} + 2 \text{Al} \longrightarrow 3 \text{Ba} + \text{Al}_2\text{O}_3$
- o.) $3 \text{O}_2 \longrightarrow 2 \text{O}_3$
- p.) $\text{SiO}_2 + 2 \text{C} \longrightarrow 2 \text{CO} + \text{Si}$
- q.) $2 \text{PbS} + 3 \text{O}_2 \longrightarrow 2 \text{PbO} + 2 \text{SO}_2$
- r.) $2 \text{HgO} \longrightarrow 2 \text{Hg} + \text{O}_2$
- s.) $\text{C}_{12}\text{H}_{22}\text{O}_{11} + 12 \text{O}_2 \longrightarrow 12 \text{CO}_2 + 11 \text{H}_2\text{O}$
- t.) $\text{C}_3\text{H}_8 + 5 \text{O}_2 \longrightarrow 3 \text{CO}_2 + 4 \text{H}_2\text{O}$
- u.) $2 \text{FeCl}_3 + 3 (\text{NH}_4)_2\text{S} \longrightarrow 6 \text{NH}_4\text{Cl} + \text{Fe}_2\text{S}_3$
- v.) $\text{P}_2\text{O}_5 + 6 \text{NaOH} \longrightarrow 2 \text{Na}_3\text{PO}_4 + 3 \text{H}_2\text{O}$
- w.) $3 \text{BaCl}_2 + \text{Al}_2(\text{SO}_4)_3 \longrightarrow 3 \text{BaSO}_4 + 2 \text{AlCl}_3$
- x.) $4 \text{NH}_3 + 5 \text{O}_2 \longrightarrow 6 \text{H}_2\text{O} + 4 \text{NO}$
- y.) $2 \text{Na}_2\text{O}_2 + 2 \text{H}_2\text{O} \longrightarrow 4 \text{NaOH} + \text{O}_2$
- z.) $\text{Sn(ClO}_3)_4 \longrightarrow \text{SnCl}_4 + 6 \text{O}_2$

Text Problems to be Completed on the Other Side of this Worksheet:

*P. 173 #45,49,53,54,56,57,58,63, 65,66,67,71

*(Although the directions may say to write an "unbalanced" equation, write a balanced equation for each.)

Unit 6 Balancing Chemical Reactions 1 Key

Kamal I. M. Al-Malah



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Basic Concepts of Chemistry Leo J. Malone, Theodore O. Dolter, 2011-12-27 The 9th edition of Malone's Basic Concepts of Chemistry provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment. New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections, assessment exercises at the end of each section, and relevant chapter problems at the end of each chapter. Every concept in the text is clearly illustrated with one or more step-by-step examples. Making it Real essays have been updated to present timely and engaging real-world applications, emphasizing the relevance of the material they are learning. This edition continues the end-of-chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter. WileyPLUS sold separately from text.

CliffsNotes AP Chemistry Angela Woodward Spangenberg, 2016-01-12 Test prep for the AP Chemistry exam with 100% brand new content that reflects recent exam changes. Addressing the major overhaul that the College Board recently made to the AP Chemistry exam, this AP Chemistry test prep guide includes completely brand new content tailored to the exam administered every May. Features of the guide include review sections of the six big ideas that the new exam focuses on: Fundamental building blocks, Molecules and interactions, Chemical reactions, Reaction rates, Thermodynamics, Chemical equilibrium. Every section includes review questions and answers. Also included in the guide are two full-length practice tests as well as a math review section and sixteen discrete laboratory exercises to prepare AP Chemistry students for the required laboratory experiments section on the exam.

A Problem-Solving Approach to Aquatic Chemistry James N. Jensen, 2023-01-05 A Problem Solving Approach to Aquatic Chemistry Enables civil and environmental engineers to understand the theory and application of aquatic equilibrium chemistry. The second edition of A Problem Solving Approach to Aquatic Chemistry provides a detailed introduction to aquatic equilibrium chemistry, calculation methods for systems at equilibrium, applications of aquatic chemistry, and chemical kinetics. The text directly addresses two required ABET program outcomes in environmental engineering chemistry: stoichiometry, equilibrium, and kinetics; and material and energy balances, fate and transport of substances in and between air, water, and soil phases. The book is very student-centered, with each chapter beginning with an introduction and ending with a summary that reviews the chapter's main points. To aid in reader comprehension, important terms are defined in context, and key ideas are summarized. Many thought-provoking discussion questions, worked examples, and end-of-chapter problems are also included. Each part of the text begins with a case study, a portion of which is addressed in each subsequent chapter, illustrating the principles of that chapter. In addition, each chapter has an Historical Note exploring connections with the people and cultures connected to topics in the text. A Problem Solving Approach to Aquatic Chemistry includes Fundamental concepts such as concentration units, thermodynamic basis of equilibrium, and manipulating equilibria. Solutions of chemical equilibrium problems including setting

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FCS Physical Science L3 Karen Morrison, 2008 Basic Concepts of Chemistry, 9e Study Guide and Solutions Manual Leo J. Malone, Theodore O. Dolter, 2012-01-03 The 9th edition of Malone's Basic Concepts of Chemistry provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections assessment exercises at the end each section and relevant chapter problems at the end of each chapter A new Math Check allows quick access to the needed basic skill The first chapter now includes brief introductions to several fundamental chemical concepts and Chapter Synthesis Problems have been added to the end of each chapter to bring key concepts into one encompassing problem Every concept in the text is clearly illustrated with one or more step by step examples Making it Real essays have been updated to present timely and engaging real world applications emphasizing the relevance of the material they are learning This edition continues the end of chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter

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Chemical Reactions and Chemical Reactors George W.

Roberts,2008-03-14 Focused on the undergraduate audience Chemical Reaction Engineering provides students with complete coverage of the fundamentals including in depth coverage of chemical kinetics By introducing heterogeneous catalysis early in the book the text gives students the knowledge they need to solve real chemistry and industrial problems An emphasis on problem solving and numerical techniques ensures students learn and practice the skills they will need later on whether for industry or graduate work

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Chemistry-vol-I Dr Vijay Sarda, Dr A C Handa, Dr K K Arora, A text book on Chemistry

College Study Skills Dianna L. Van Blerkom, 2000 This comprehensive text workbook focuses on getting students actively involved in the learning process and in learning how to learn Rather than presenting one prescriptive approach students take a learning style inventory and are then exposed to strategies that complement their personal learning styles

Physical Chemistry for Engineering and Applied Sciences Frank R. Foulkes, 2012-09-12 Physical Chemistry for Engineering and Applied Sciences is the product of over 30 years of teaching first year Physical Chemistry as part of the Faculty of Applied Science and Engineering at the University of Toronto Designed to be as rigorous as compatible with a first year student s ability to understand the text presents detailed step by step derivations of the equations that permit the student to follow the underlying logic and of equal importance to appreciate any simplifying assumptions made or mathematical tricks employed In addition to the 600 exercises and end of chapter problems the text is rich in worked non trivial examples many of which are designed to be inspiring and thought provoking Step by step derivation of all equations enables the student to smoothly follow the derivation by sight and can be understood relatively easily by students with moderate skills and backgrounds in mathematics Clear and accessible Physical Chemistry for Engineering and Applied Sciences includes The answers to all of the 112 worked examples 99 exercises following many of the worked examples and 496 end of chapter problems Topics not normally seen in introductory physical chemistry textbooks ionic reaction rates activities and activity coefficients or not regularly explained in much detail electrochemistry chemical kinetics with an eye on industrial applications Special appendices that provide detailed explanations of basic integration and

natural logarithms for students lacking a background in integral calculus An in depth chapter on electrochemistry in which activities and activity coefficients are used extensively as required for accurate calculations

Handbook on Material and Energy Balance Calculations in Material Processing Arthur E. Morris, Gordon Geiger, H. Alan Fine, 2012-01-03 Lately there has been a renewed push to minimize the waste of materials and energy that accompany the production and processing of various materials This third edition of this reference emphasizes the fundamental principles of the conservation of mass and energy and their consequences as they relate to materials and energy New to this edition are numerous worked examples illustrating conventional and novel problem solving techniques in applications such as semiconductor processing environmental engineering the production and processing of advanced and exotic materials for aerospace electronic and structural applications

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