

Stoichiometry "End of Chapter" IN CLASS Practice Assessment

Background Info -

- Why are the coefficients in a balanced equation necessary for stoichiometry?
Coefficients are necessary for establishing a molar ratio to convert between moles of different substances.

Skill 1 - Using a balanced equation, calculate using mole to mole ratios. Use a balanced equation, given the grams of reactant, calc. the grams of product.

- Consider the following unbalanced equation:



There are 5.56g of manganese(IV) oxide present, calculate the grams of chlorine gas.

$$5.00\text{g MnO}_2 \times \frac{1\text{mol MnO}_2}{86.94\text{g MnO}_2} \times \frac{1\text{mol Cl}_2}{1\text{mol MnO}_2} \times \frac{70.90\text{g Cl}_2}{1\text{mol Cl}_2} = 4.00\text{g Cl}_2$$

Skill 2 - After calculating theoretical yield in grams, calculate the percent yield.

- You are assigned to complete a lab. Before the reaction is complete, you calculate that a reaction should produce 4.0g Cu(s). After the lab is over you find that 2.8g Cu(s) was collected. What would the percent yield be for that reaction completed within the lab?

$$\% \text{yield} = \left[\frac{\text{actual}}{\text{theoretical}} \right] \times 100 \\ = \left[\frac{2.8\text{g}}{4.0\text{g}} \right] \times 100 = 70\%$$

Skill 3 - Given experimental data, determine the limiting reactant and the excess reactant.

- Calcium carbonate combines with gaseous hydrogen chloride to produce solid calcium chloride, gaseous carbon dioxide, and water vapor. What is the percent yield if 155g of calcium carbonate combines with 250.0g of hydrogen chloride and only 142.0g of calcium chloride is actually obtained within the lab setting? Complete and show work for all 3 steps of a stoichiometry problem.



$$155\text{g CaCO}_3 \times \frac{1\text{mol CaCO}_3}{100.09\text{g CaCO}_3} = 1.55\text{mol CaCO}_3 \quad \text{L.R.} \\ 1.55\text{mol CaCO}_3 \times \frac{1\text{mol CaCl}_2}{1\text{mol CaCO}_3} = 1.55\text{mol CaCl}_2$$

$$250\text{g HCl} \times \frac{1\text{mol HCl}}{36.458\text{g HCl}} = 6.857\text{mol HCl} \quad \text{E.R.} \\ 6.857\text{mol HCl} \times \frac{1\text{mol CaCl}_2}{2\text{mol HCl}} = 3.43\text{mol CaCl}_2$$

$$1.55\text{mol CaCO}_3 \times \left(\frac{1\text{mol CaCl}_2}{1\text{mol CaCO}_3} \right) \times \left(\frac{110.98\text{g CaCl}_2}{1\text{mol CaCl}_2} \right) = 172.02\text{g}$$

$$\% \text{yield} = \left[\frac{\text{actual}}{\text{theoretical}} \right] \times 100 \\ = \left[\frac{142.0\text{g}}{172.02\text{g}} \right] \times 100 = 82.55\%$$

Unit 7 Stoichiometry Review

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Reviews of Physiology, Biochemistry and Pharmacology, 2007-05-02 H. Wegele, L. M. Iler and J. Buchner. *Hsp70 and Hsp90: A Relay Team for Protein Folding*. R. Schlein. *The Early Stages of the Intracellular Transport of Membrane Proteins: Clinical and Pharmacological Implications*. L. Schild. *The Epithelial Sodium Channel: From Molecule to Disease*.

Merrill Chemistry Robert C. Smoot, Smoot, Richard G. Smith, Jack Price, 1998. *Nuclear Science Abstracts*, 1972.

Annual Review of Biophysics and Biophysical Chemistry Donald M. Engelman, 1987. *Resources in Education*, 1980-04.

Annual Review of Materials Science Robert A. Huggins, 1992. Provides abstracts and full text to articles on materials science.

Product and Process Design Jan Harmsen, André B. de Haan, Pieter L. J. Swinkels, 2018-05-22. *Product and Process Design: Driving Innovation* is a comprehensive textbook for students and industrial professionals. It treats the combined design of innovative products and their innovative manufacturing processes, providing specific methods for BSc, MSc, PDEng, and PhD courses. Students, industrial innovators, and managers are guided through all design steps in all innovation stages: discovery, concept, feasibility, development, detailed engineering, and implementation, to successfully obtain novel products and their novel processes. The authors' decades of innovation experience in industry as well as in teaching BSc, MSc, and post-academic product and process design courses, thereby including the latest design publications, culminate in this book.

Annual Review of Biophysics and Biophysical Chemistry, 1991. *Annual Review of Biochemistry* E. E. Snell, 1983.

Energy Research Abstracts, 1989. **Annual Review of Biophysics and Bioengineering** L. J. Mullins, 1975.

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