

Unit 2 Worksheet: St. Michael's Concentration

Part 2: Describing Molar Concentrations

1. Sodium chloride was dissolved in water to produce a 1.5M solution. Suppose that this same solution had been added to the flask solution. This would be incorrect. **Just this solution.**
Then, calculate the molarities of the two solutions. The molarity of the solution is 1.5M. The molarity of the solution is 1.5M.
2. A 10.0 g solution of potassium chloride is dissolved in enough water to make 100.0 mL of solution. Calculate the molar concentration of this potassium chloride.

$$10.0 \text{ g} \times \frac{1 \text{ mol}}{74.55 \text{ g}} = 0.134 \text{ mol}$$

$$\text{Molarity} = \frac{0.134 \text{ mol}}{0.100 \text{ L}} = 1.34 \text{ M}$$

3. Find the molarity of a solution made from 25.0 g of CaCl_2 dissolved in enough water to make 0.50 L.

$$25.0 \text{ g} \times \frac{1 \text{ mol}}{110.98 \text{ g}} = 0.225 \text{ mol}$$

$$\text{Molarity} = \frac{0.225 \text{ mol}}{0.50 \text{ L}} = 0.45 \text{ M}$$

4. An unknown ionic solution's "molarity" is defined by preparing a 1.00 g of solution sample and measuring the volume of solution in mL. Calculate the molarity of the solution in the solution.

$$1.00 \text{ g} \times \frac{1 \text{ mol}}{74.55 \text{ g}} = 0.0134 \text{ mol}$$

$$\text{Molarity} = \frac{0.0134 \text{ mol}}{0.100 \text{ L}} = 0.134 \text{ M}$$

5. A 10.0 g solution of potassium chloride is dissolved in enough water to make 100.0 mL of solution. Calculate the molar concentration of this potassium chloride.

$$10.0 \text{ g} \times \frac{1 \text{ mol}}{74.55 \text{ g}} = 0.134 \text{ mol} \quad \frac{0.134 \text{ mol}}{0.100 \text{ L}} = 1.34 \text{ M}$$

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Table of Contents Unit 1worksheet 3 Molar Concentration

1. Understanding the eBook Unit 1worksheet 3 Molar Concentration
 - The Rise of Digital Reading Unit 1worksheet 3 Molar Concentration
 - Advantages of eBooks Over Traditional Books
2. Identifying Unit 1worksheet 3 Molar Concentration
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Unit 1worksheet 3 Molar Concentration
 - User-Friendly Interface
4. Exploring eBook Recommendations from Unit 1worksheet 3 Molar Concentration
 - Personalized Recommendations
 - Unit 1worksheet 3 Molar Concentration User Reviews and Ratings
 - Unit 1worksheet 3 Molar Concentration and Bestseller Lists
5. Accessing Unit 1worksheet 3 Molar Concentration Free and Paid eBooks
 - Unit 1worksheet 3 Molar Concentration Public Domain eBooks

- Unit 1 worksheet 3 Molar Concentration eBook Subscription Services
- Unit 1 worksheet 3 Molar Concentration Budget-Friendly Options
- 6. Navigating Unit 1 worksheet 3 Molar Concentration eBook Formats
 - ePub, PDF, MOBI, and More
 - Unit 1 worksheet 3 Molar Concentration Compatibility with Devices
 - Unit 1 worksheet 3 Molar Concentration Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Unit 1 worksheet 3 Molar Concentration
 - Highlighting and Note-Taking Unit 1 worksheet 3 Molar Concentration
 - Interactive Elements Unit 1 worksheet 3 Molar Concentration
- 8. Staying Engaged with Unit 1 worksheet 3 Molar Concentration
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Unit 1 worksheet 3 Molar Concentration
- 9. Balancing eBooks and Physical Books Unit 1 worksheet 3 Molar Concentration
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Unit 1 worksheet 3 Molar Concentration
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Unit 1 worksheet 3 Molar Concentration
 - Setting Reading Goals Unit 1 worksheet 3 Molar Concentration
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Unit 1 worksheet 3 Molar Concentration
 - Fact-Checking eBook Content of Unit 1 worksheet 3 Molar Concentration
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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