

## Unit 3 Worksheet 4 – Quantitative Energy Problems Part 2

### Energy constants (H<sub>2</sub>O)

|            |                                                        |
|------------|--------------------------------------------------------|
| 334 J/g    | Heat of fusion (melting or freezing) $H_f$             |
| 2260 J/g   | Heat of vaporization (evaporating or condensing) $H_v$ |
| 2.1 J/g°C  | Heat capacity (c) of solid water                       |
| 4.18 J/g°C | Heat capacity (c) of liquid water                      |

For each of the problems sketch a warming or cooling curve to help you decide which equation(s) to use to solve the problem. Keep a reasonable number of sig figs in your answers.

1. How much energy must be absorbed by a 150 g sample of ice at 0.0 °C that melts and then warms to 25.0°C?
2. Suppose in the Icy Hot lab that the burner transfers 325 kJ of energy to 450 g of liquid water at 20.°C. What mass of the water would be boiled away?
3. A 12oz can of soft drink (assume  $m = 340$  g) at 25°C is placed in a freezer where the temperature is - 12 °C. How much energy must be removed from the soft drink for it to reach this temperature?

# Unit 3 4 Quantitative Energy Problems Answers

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