

Chemistry Unit 8

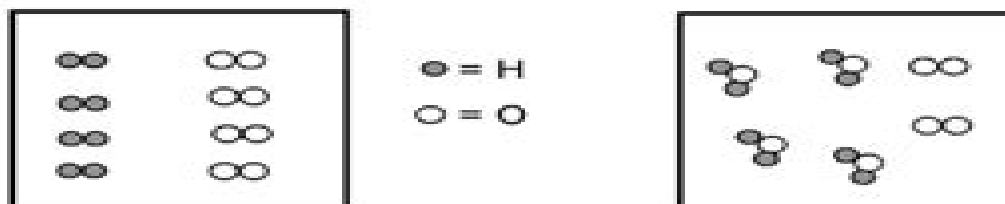
Worksheet 3: Adjusting to Reality - Limiting Reactant

1. Write the balanced equation for the reaction between hydrogen and oxygen.

Balanced Equation: 2 $\text{H}_{2(g)}$ + 1 $\text{O}_{2(g)}$ $\rightarrow$ 2 $\text{H}_2\text{O}_{(g)}$

Suppose that 4 molecules of hydrogen gas and 4 molecules of oxygen gas react to form water.

Make a drawing that represents the reaction container before and after the reaction.



Before

After

4 molecules How many molecules of water can be produced?

oxygen Which reactant is in excess? Why? There are leftover molecules

2 molecules How many molecules of excess reactant are there?

Construct a Before-Change-After Table for this reactant mixture:

In this table the numbers refer to molecules rather than moles.

Equation: 2 $\text{H}_{2(g)}$ + 1 $\text{O}_{2(g)}$ $\rightarrow$ 2 $\text{H}_2\text{O}_{(g)}$

Before 4 4 0

Change -4 -2 +4

After 0 2 4

According to the table you just made,

4 molecules How many molecules of water can be produced?

oxygen Which reactant is in excess? Why? 2 molecules are left-over "after" the reaction

2 molecules How many molecules of excess reactant are there?

Based on your two methods of analysis above, what determines how much product can be made from a particular reactant mix?

The determining factor is which of the reactants runs out first. This determines how much product can be made, even if another reactant is in excess.

Unit 7 3 Limiting Reactant Answers

Yicheng Fang



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