

Unit 1 Chemistry Study Guide

History of the Atom:

Democritus (450 BCE):

- Matter is not infinitely divisible.
- There are single particles that cannot be divided further, called **atomos**.

Aristotle (380 BCE):

- Matter is infinitely divisible.
- Matter is called **hyle**.

Antoine Lavoisier (1774):

- **Law of Conservation of Mass:** Mass cannot be created or destroyed.

Joseph Proust (1779):

- **Law of Definite Proportions:** All samples of a pure compound contain the same elements in the same proportion by mass (no matter quantity).

John Dalton (1803 – 1808):

- Observed work of Lavoisier and Proust
- Led Dalton to develop the first **atomic theory**.
- **Dalton's Atomic Theory:**
 - o All matter is composed of extremely small particles called atoms.
 - o All atoms of a given element are identical, having the same size, mass, and chemical properties. The atoms of one element are different from the atoms of all other elements.
 - o Atoms cannot be subdivided, created, or destroyed.
 - o Atoms of different elements can combine in simple whole number ratios to form chemical compounds.
 - o In chemical reactions, atoms are combined, separated or rearranged.
- **Law of Multiple Proportions:** Elements can combine in different proportions. However, the different compounds will have different properties but the ratio of atoms is always a whole number.

Billiard Ball Model:



atom

J.J. Thompson (1906):

- **Cathode Ray Experiment –**
 - o Part 1: Thompson ran electricity through a cathode ray tube and when he put the negatively charged end of a magnet near the tube, he saw the ray of electricity

Unit 8 Study Guide Chemistry Answers

Ted Lister, Janet Renshaw



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Study Guide and Solutions Manual for McMurry's Organic Chemistry, Fifth Edition Susan McMurry, 2000

Provides answers and explanations to all in text and end of chapter exercises Also includes summaries of name reactions functional group synthesis and reactions lists of reagents and abbreviations and articles on topics ranging from infrared absorption frequencies to the Nobel Prize winners in Chemistry This edition now includes all new artwork expanded in text problems summary quizzes approximately every three chapters more detailed explanations in solutions and chapter outlines

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Empowering Formal and Informal Leadership While Maintaining Teacher Identity Zugelder, Bryan S., 2021-05-07 Teacher

leadership remains at the forefront of conversations in teacher education with discussions on recruitment retention and effectiveness Teachers are at the core of schooling and the roles they assume and types of leadership they engage in are multi dimensional Teacher leadership comes in many shapes and definitions In this sense both opportunities and challenges exist in teacher leadership While national competencies continue to define dispositional and knowledge base for teacher leaders there is still work to be done to define and add to the body of scholarship on this topic Teacher leadership opportunities provide development for teachers as a key retention strategy however role ambiguity presents challenges in how to empower teachers for formal and informal roles of leadership while maintaining the teacher identity Empowering Formal and Informal Leadership While Maintaining Teacher Identity provides a comprehensive look at the opportunities and challenges of teacher leadership drawing on research and practice that add to the body of knowledge for teacher leadership which is a niche in education that differs from educational administration Highlighted topics within this book include models for teacher leadership teacher education types of leadership roles the development and retention of teacher leaders and leadership across different types of school districts This book is ideally designed for inservice and preservice teachers administrators teacher educators practitioners researchers academicians and students interested in teacher leadership roles and the way in which teacher identity is maintained in relation to these other leadership positions **CliffsNotes AP**

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