

CHEM 101 Quiz #6 Chapter 6 ANSWERS**1. (4 points)**

The specific heat (s) of iron (Fe) is 0.40 J/(g K) . Calculate the heat effect q when 20 g of iron is heated from 100°C to 200°C . Include the sign!

$$q = m \cdot s \cdot \Delta T \text{ for } \Delta T \Delta K = \Delta^\circ\text{C} \quad q = 20 \times 0.40 \times (200 - 100) = +800 \text{ J}$$

You must include the units and sign in the answer!

2. (4 points)

The Molar Mass of Mg = 24.0 g/mol .

Calculate the enthalpy change when 2.40 g Mg(s) reacts with excess oxygen. Include the sign!

$$2.40 \text{ g Mg} = 2.40/24.0 = 0.1 \text{ mol} \rightarrow \Delta H = 0.1(\text{mol}) \times (-600)(\text{kJ/mol}) = -60 \text{ kJ}$$

Again, your answer must include both the sign and the units.

3. (2 points) Which one of the following four reactions is a *formation reaction*? Circle the correct answer.



Look up the definition of a formation reaction in Chang or in your lecture notes: the compound is formed from its elements, and the elements must be as they naturally occur so e.g. Cl_2 not Cl , and H_2 not H !

CHEM 101 Section L52 Quiz #6**1. (4 points)**

The specific heat (s) of copper (Cu) is 0.38 J/(g K) . Calculate the heat effect q when 10 g of copper is heated from 100°C to 300°C . Include the sign!

$$q = m \cdot s \cdot \Delta T \text{ for } \Delta T \Delta K = \Delta^\circ\text{C} \quad q = 10 \times 0.38 \times (300 - 100) = +760 \text{ J}$$

You must include the units and sign in the answer!

2. (4 points)

The Molar Mass of Ca = 40.0 g/mol .

Calculate the enthalpy change ΔH when 8.0 g Ca(s) reacts with excess oxygen. Include the sign!!

$$8.0 \text{ g Ca} = 8.0/40 = 0.2 \text{ mol} \rightarrow \Delta H = 0.2(\text{mol}) \times (-650)(\text{kJ/mol}) = -130 \text{ kJ}$$

Again, your answer must include both the sign and the units.

3. (2 points) Which one of the following reactions is a *formation reaction*? Circle the correct answer.



Unit 6 6 Chemistry Answers

**D M Vasudevan, Sreekumari S, Kannan
Vaidyanathan**



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Computational Chemistry Errol G. Lewars, 2016-09-20 This is the third edition of the successful text reference book that covers computational chemistry It features changes to the presentation of key concepts and includes revised and new material with several expanded exercises at various levels such as harder questions for those ready to be tested in greater depth this aspect is absent from other textbooks in the field Although introductory and assuming no prior knowledge of computational chemistry it covers the essential aspects of the subject There are several introductory textbooks on computational chemistry this one is as in its previous editions a unique textbook in the field with copious exercises and questions and solutions with discussions Noteworthy is the fact that it is the only book at the introductory level that shows in detail yet clearly how matrices are used in one important aspect of computational chemistry It also serves as an essential guide for researchers and as a reference book

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