ssessment #3 Review	Name:
ou must show all work for	credit.
ie Date:	(No late packets will be accepted)
I. Balance Equation	ns.
Balance the equations	F II True
 Mg + ∠ H₂O → M; 	$g(OH)_2 + H_2$
 Mg + ZNaBr → N 	MgBr ₂ + ZNa
3. C ₃ H ₈ + 5 O ₂ →3 C	
4. 4 P + 5 O ₂ → 2 P ₂ O ₅	1 1 1 1 CO 1 7 Ec/OH)
5. Fe ₂ (SO ₄) ₃ + 6 KOH	I →3 K₂SO₄ + 2 Fe(OH)₃
 Fe₂O₃ + 3 H₂ → F 	7
7 How many moles of	hydrogen gas react with 1 mole of Fe₂O₃ in reaction #6?
7. 11011 1112117	
rite the balanced equation	n for problems # 8 and 9:
Hydrogen gas and n	nitrogen monoxide combine to form water and nitrogen gas
C Man and make a	phate and calcium chloride are mixed they form calcium phosphate and
sodium chloride	2 Nas 704 13 Cacle - Cas (PON) = + 8 Na C
Sodium chloride	< Nastry 13 cacks - castronie actions
10 Using the following b	balanced equation, prove (or explain) the Law of Conservation of Mass.
$N_2 + 3H_2 \rightarrow 2NH_3$	
	Z N = lety Z = 1 ; [w]
	su no plants created or lost
The second second second second	equation shown in problem #10, explain what the coefficients represent.
11. Using the balanced	equation shown in problem #10, explain what the coemoleria represents
12. Using the balanced	equation shown in problem #10, list the products of the reaction.
	UKZ
13. Using the balanced	equation shown in problem #10, list the reactants of the reaction.
	1 12
II. Types of Reacti	ions
irections: Identify the type	e of reaction.
= synthesis	SD = single displacement C= combustion DD = double displacement
= decomposition	DD = double displacement
14. HCI + MgS	$O_4 \rightarrow MgCl_2 + H_2SO_4$
15. Cu₂(CO₃)₃ →	Cu + CO ₂ + O ₂
16 C ₂ H ₆ + O ₂	→ CO ₂ + H ₂ O
	→ CO ₂ + H ₂ O
16. $C_2H_6 + O_2$	\rightarrow CO ₂ + H ₂ O P ₂ O ₅
16. $C_2H_6 + O_2$	\rightarrow CO ₂ + H ₂ O P ₂ O ₅ CaCl ₂ \rightarrow NaCl + Ca(NO ₃) ₂

Unit 6 Test Chemistry

Padhraic Smyth

Unit 6 Test Chemistry:	
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This book delves into Unit 6 Test Chemistry. Unit 6 Test Chemistry is a vital topic that needs to be grasped by everyone, ranging from students and scholars to the general public. This book will furnish comprehensive and in-depth insights into Unit 6 Test Chemistry, encompassing both the fundamentals and more intricate discussions.

- 1. This book is structured into several chapters, namely:
 - Chapter 1: Introduction to Unit 6 Test Chemistry
 - Chapter 2: Essential Elements of Unit 6 Test Chemistry
 - Chapter 3: Unit 6 Test Chemistry in Everyday Life
 - Chapter 4: Unit 6 Test Chemistry in Specific Contexts
 - ∘ Chapter 5: Conclusion
- 2. In chapter 1, this book will provide an overview of Unit 6 Test Chemistry. The first chapter will explore what Unit 6 Test Chemistry is, why Unit 6 Test Chemistry is vital, and how to effectively learn about Unit 6 Test Chemistry.
- 3. In chapter 2, this book will delve into the foundational concepts of Unit 6 Test Chemistry. This chapter will elucidate the essential principles that must be understood to grasp Unit 6 Test Chemistry in its entirety.
- 4. In chapter 3, this book will examine the practical applications of Unit 6 Test Chemistry in daily life. This chapter will showcase real-world examples of how Unit 6 Test Chemistry can be effectively utilized in everyday scenarios.
- 5. In chapter 4, the author will scrutinize the relevance of Unit 6 Test Chemistry in specific contexts. The fourth chapter will explore how Unit 6 Test Chemistry is applied in specialized fields, such as education, business, and technology.
- 6. In chapter 5, the author will draw a conclusion about Unit 6 Test Chemistry. The final chapter will summarize the key points that have been discussed throughout the book.
 - The book is crafted in an easy-to-understand language and is complemented by engaging illustrations. It is highly recommended for anyone seeking to gain a comprehensive understanding of Unit 6 Test Chemistry.

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