

Balancing Equations Worksheet

- 1) $\underline{\hspace{1cm}} \text{H}_3\text{PO}_4 + \underline{\hspace{1cm}} \text{KOH} \rightarrow \underline{\hspace{1cm}} \text{K}_3\text{PO}_4 + \underline{\hspace{1cm}} \text{H}_2\text{O}$
- 2) $\underline{\hspace{1cm}} \text{K} + \underline{\hspace{1cm}} \text{B}_2\text{O}_3 \rightarrow \underline{\hspace{1cm}} \text{K}_2\text{O} + \underline{\hspace{1cm}} \text{B}$
- 3) $\underline{\hspace{1cm}} \text{HCl} + \underline{\hspace{1cm}} \text{NaOH} \rightarrow \underline{\hspace{1cm}} \text{NaCl} + \underline{\hspace{1cm}} \text{H}_2\text{O}$
- 4) $\underline{\hspace{1cm}} \text{Na} + \underline{\hspace{1cm}} \text{NaNO}_3 \rightarrow \underline{\hspace{1cm}} \text{Na}_2\text{O} + \underline{\hspace{1cm}} \text{N}_2$
- 5) $\underline{\hspace{1cm}} \text{C} + \underline{\hspace{1cm}} \text{S}_8 \rightarrow \underline{\hspace{1cm}} \text{CS}_2$
- 6) $\underline{\hspace{1cm}} \text{Na} + \underline{\hspace{1cm}} \text{O}_2 \rightarrow \underline{\hspace{1cm}} \text{Na}_2\text{O}$
- 7) $\underline{\hspace{1cm}} \text{N}_2 + \underline{\hspace{1cm}} \text{O}_2 \rightarrow \underline{\hspace{1cm}} \text{N}_2\text{O}_5$
- 8) $\underline{\hspace{1cm}} \text{H}_3\text{PO}_4 + \underline{\hspace{1cm}} \text{Mg}(\text{OH})_2 \rightarrow \underline{\hspace{1cm}} \text{Mg}_3(\text{PO}_4)_2 + \underline{\hspace{1cm}} \text{H}_2\text{O}$
- 9) $\underline{\hspace{1cm}} \text{NaOH} + \underline{\hspace{1cm}} \text{H}_2\text{CO}_3 \rightarrow \underline{\hspace{1cm}} \text{Na}_2\text{CO}_3 + \underline{\hspace{1cm}} \text{H}_2\text{O}$
- 10) $\underline{\hspace{1cm}} \text{KOH} + \underline{\hspace{1cm}} \text{HBr} \rightarrow \underline{\hspace{1cm}} \text{KBr} + \underline{\hspace{1cm}} \text{H}_2\text{O}$
- 11) $\underline{\hspace{1cm}} \text{Na} + \underline{\hspace{1cm}} \text{O}_2 \rightarrow \underline{\hspace{1cm}} \text{Na}_2\text{O}$
- 12) $\underline{\hspace{1cm}} \text{Al}(\text{OH})_3 + \underline{\hspace{1cm}} \text{H}_2\text{CO}_3 \rightarrow \underline{\hspace{1cm}} \text{Al}_2(\text{CO}_3)_3 + \underline{\hspace{1cm}} \text{H}_2\text{O}$
- 13) $\underline{\hspace{1cm}} \text{Al} + \underline{\hspace{1cm}} \text{S}_8 \rightarrow \underline{\hspace{1cm}} \text{Al}_2\text{S}_3$
- 14) $\underline{\hspace{1cm}} \text{Cs} + \underline{\hspace{1cm}} \text{N}_2 \rightarrow \underline{\hspace{1cm}} \text{Cs}_3\text{N}$
- 15) $\underline{\hspace{1cm}} \text{Mg} + \underline{\hspace{1cm}} \text{Cl}_2 \rightarrow \underline{\hspace{1cm}} \text{MgCl}_2$
- 16) $\underline{\hspace{1cm}} \text{Rb} + \underline{\hspace{1cm}} \text{RbNO}_3 \rightarrow \underline{\hspace{1cm}} \text{Rb}_2\text{O} + \underline{\hspace{1cm}} \text{N}_2$
- 17) $\underline{\hspace{1cm}} \text{C}_6\text{H}_6 + \underline{\hspace{1cm}} \text{O}_2 \rightarrow \underline{\hspace{1cm}} \text{CO}_2 + \underline{\hspace{1cm}} \text{H}_2\text{O}$
- 18) $\underline{\hspace{1cm}} \text{N}_2 + \underline{\hspace{1cm}} \text{H}_2 \rightarrow \underline{\hspace{1cm}} \text{NH}_3$
- 19) $\underline{\hspace{1cm}} \text{C}_{10}\text{H}_{22} + \underline{\hspace{1cm}} \text{O}_2 \rightarrow \underline{\hspace{1cm}} \text{CO}_2 + \underline{\hspace{1cm}} \text{H}_2\text{O}$
- 20) $\underline{\hspace{1cm}} \text{Al}(\text{OH})_3 + \underline{\hspace{1cm}} \text{HBr} \rightarrow \underline{\hspace{1cm}} \text{AlBr}_3 + \underline{\hspace{1cm}} \text{H}_2\text{O}$
- 21) $\underline{\hspace{1cm}} \text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_3 + \underline{\hspace{1cm}} \text{O}_2 \rightarrow \underline{\hspace{1cm}} \text{CO}_2 + \underline{\hspace{1cm}} \text{H}_2\text{O}$
- 22) $\underline{\hspace{1cm}} \text{C}_7\text{H}_8 + \underline{\hspace{1cm}} \text{O}_2 \rightarrow \underline{\hspace{1cm}} \text{CO}_2 + \underline{\hspace{1cm}} \text{H}_2\text{O}$
- 23) $\underline{\hspace{1cm}} \text{Li} + \underline{\hspace{1cm}} \text{AlCl}_3 \rightarrow \underline{\hspace{1cm}} \text{LiCl} + \underline{\hspace{1cm}} \text{Al}$
- 24) $\underline{\hspace{1cm}} \text{C}_2\text{H}_6 + \underline{\hspace{1cm}} \text{O}_2 \rightarrow \underline{\hspace{1cm}} \text{CO}_2 + \underline{\hspace{1cm}} \text{H}_2\text{O}$
- 25) $\underline{\hspace{1cm}} \text{NH}_4\text{OH} + \underline{\hspace{1cm}} \text{H}_3\text{PO}_4 \rightarrow \underline{\hspace{1cm}} (\text{NH}_4)_3\text{PO}_4 + \underline{\hspace{1cm}} \text{H}_2\text{O}$
- 26) $\underline{\hspace{1cm}} \text{Rb} + \underline{\hspace{1cm}} \text{P} \rightarrow \underline{\hspace{1cm}} \text{Rb}_3\text{P}$
- 27) $\underline{\hspace{1cm}} \text{CH}_4 + \underline{\hspace{1cm}} \text{O}_2 \rightarrow \underline{\hspace{1cm}} \text{CO}_2 + \underline{\hspace{1cm}} \text{H}_2\text{O}$
- 28) $\underline{\hspace{1cm}} \text{Al}(\text{OH})_3 + \underline{\hspace{1cm}} \text{H}_2\text{SO}_4 \rightarrow \underline{\hspace{1cm}} \text{Al}_2(\text{SO}_4)_3 + \underline{\hspace{1cm}} \text{H}_2\text{O}$
- 29) $\underline{\hspace{1cm}} \text{Na} + \underline{\hspace{1cm}} \text{Cl}_2 \rightarrow \underline{\hspace{1cm}} \text{NaCl}$
- 30) $\underline{\hspace{1cm}} \text{Rb} + \underline{\hspace{1cm}} \text{S}_8 \rightarrow \underline{\hspace{1cm}} \text{Rb}_2\text{S}$
- 31) $\underline{\hspace{1cm}} \text{H}_3\text{PO}_4 + \underline{\hspace{1cm}} \text{Ca}(\text{OH})_2 \rightarrow \underline{\hspace{1cm}} \text{Ca}_3(\text{PO}_4)_2 + \underline{\hspace{1cm}} \text{H}_2\text{O}$
- 32) $\underline{\hspace{1cm}} \text{NH}_3 + \underline{\hspace{1cm}} \text{HCl} \rightarrow \underline{\hspace{1cm}} \text{NH}_4\text{Cl}$
- 33) $\underline{\hspace{1cm}} \text{Li} + \underline{\hspace{1cm}} \text{H}_2\text{O} \rightarrow \underline{\hspace{1cm}} \text{LiOH} + \underline{\hspace{1cm}} \text{H}_2$
- 34) $\underline{\hspace{1cm}} \text{Ca}_3(\text{PO}_4)_2 + \underline{\hspace{1cm}} \text{SiO}_2 + \underline{\hspace{1cm}} \text{C} \rightarrow \underline{\hspace{1cm}} \text{CaSiO}_3 + \underline{\hspace{1cm}} \text{CO} + \underline{\hspace{1cm}} \text{P}$
- 35) $\underline{\hspace{1cm}} \text{NH}_3 + \underline{\hspace{1cm}} \text{O}_2 \rightarrow \underline{\hspace{1cm}} \text{N}_2 + \underline{\hspace{1cm}} \text{H}_2\text{O}$
- 36) $\underline{\hspace{1cm}} \text{FeS}_2 + \underline{\hspace{1cm}} \text{O}_2 \rightarrow \underline{\hspace{1cm}} \text{Fe}_2\text{O}_3 + \underline{\hspace{1cm}} \text{SO}_2$
- 37) $\underline{\hspace{1cm}} \text{C} + \underline{\hspace{1cm}} \text{SO}_2 \rightarrow \underline{\hspace{1cm}} \text{CS}_2 + \underline{\hspace{1cm}} \text{CO}$

Unit 7 Balancing Chemical Reactions Key

Darrell D. Ebbing, Albert E. Richardson

A red circular graphic with a gradient, appearing as a partial circle or a stylized arrow pointing to the right, located to the right of the authors' names.

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