

- 2)  $\underline{1} \text{ HBr} + \underline{1} \text{ KHCO}_3 \rightarrow \underline{1} \text{ H}_2\text{O} + \underline{1} \text{ KBr} + \underline{1} \text{ CO}_2$
- 3)  $\underline{2} \text{ GaBr}_3 + \underline{3} \text{ Na}_2\text{SO}_3 \rightarrow \underline{1} \text{ Ga}_2(\text{SO}_3)_3 + \underline{6} \text{ NaBr}$
- 4)  $\underline{3} \text{ SnO} + \underline{2} \text{ NF}_3 \rightarrow \underline{3} \text{ SnF}_2 + \underline{1} \text{ N}_2\text{O}_3$

Using the equation from problem 2 above, answer the following questions:

- 5) If I do this reaction with 30 grams of HBr, how many grams of carbon dioxide will be formed? (2 points)

- 6) Using the following equation:



How many grams of sodium sulfate will be formed if you start with 150 grams of sodium hydroxide and you have an excess of sulfuric acid? (2 points)

- 7) Using the following equation:



How many grams of lithium nitrate will be needed to make 350 grams of lithium sulfate, assuming that you have an adequate amount of lead (IV) sulfate to do the reaction? (2 points)

# Unit 8 Stoichiometry Test Practice

**Dr. K. G. Ojha**



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