11 BALANCING CHEMICAL EQUATIONS

Directions:. Label decomposition, synthesis (combination), single displacement, double displacement (or double replacement), combustion reaction or Neutralization (acid/base).

Balance the Chemical Equation	Reaction Type
1. 2 K + 1 Cl ₂ → 2 KCl	COMBINATION
2. $_SO_2 + _O_2 \rightarrow _SO_3$	S
3KCIO ₃ →KCI +O ₂	<u> </u>
4CO ₂ + _C → _CO	e
5H ₂ +O ₂ →H ₂ O	= -
6MgCO ₃ →MgO +O ₂	-
7. $\underline{\hspace{0.1cm}}$	<u> </u>
8. $_{\text{Fe}_2\text{O}_3} + _{\text{Al}} \rightarrow _{\text{Fe}} + _{\text{Al}_2\text{O}_3}$	S <u>. </u>
9. $_PbCl_2 + _Li_2SO_4 \rightarrow _PbSO_4 + _LiCI$	82
10. $_NaOH + \H_2SO_4 \rightarrow \Na_2SO_4 + \H-OH$	
12Pb(OH) ₂ +HCl →PbCl ₂ +H-OH	
13Na ₂ CO ₃ +HCl →H ₂ CO ₃ +NaCl	<u> </u>
14AgNO ₃ +Cu →Ag +Cu(NO ₃) ₂	-
15. $_CH_4 + _O_2 \rightarrow _CO_2 + _H_2O$	S
16. $C_2H_6 + O_2 \rightarrow CO_2 + H_2O$	-
17. $_C_3H_8 + _O_2 \rightarrow _CO_2 + _H_2O$	
18. $_C_4H_{10} + _O_2 \rightarrow _CO_2 + _H_2O$	
19. $_C_5H_{12} + _O_2 \rightarrow _CO_2 + _H_2O$	·

Bonus: Don't unless given permission. Use decimal method if needed, for example "1.5 O2."

20.
$$C_7H_5N_3O_6 + O_2 \rightarrow CO_2 + H_2O + N_2$$

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