

Covalent Bonds

Covalent Bonds

are the forces that hold atoms together in their compounds

Valence Electrons

are the s and p electrons in the highest occupied energy level. They are involved in chemical bonds.

Four main types of bonds:

ionic, covalent, metallic, & hydrogen

Valence Electrons

are s electrons in their valence shells, which is why they will react. Some bonds under certain circumstances.

A chemical bond is an interaction between atoms or ions that results in a lower potential energy of the system. This lower potential energy means the system more stable. A chemical bond of the system now has the lower energy.

Covalent Bonds

Covalent bonds are attractive forces that are produced by the sharing of electrons.

Substances containing covalent bonds form molecules. Covalent bonds are between 2 non-metals and hydrogen.

eg: water



Polar Covalent vs. Non-Polar Covalent

Electronegativity greater than 1.7 = ionic bond is formed

Less than 1.7 = covalent & it can be polar or non-polar depending on electronegativity

Polar Covalent = unequal Non-Polar Covalent = equal

EXAMPLES: non-polar and non-polar compounds:

hydrocarbons like

hydrocarbons like

non-polar covalent structures are polar covalent things. Water is polar covalent and it

is non-polar covalent and that is why they don't mix.

Polar bonds compounds dissolve in polar covalent substances

Workbook Pages 8 1 Covalent Bonds Guided Pages Answers

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Workbook Pages 8 1 Covalent Bonds Guided Pages Answers Book Review: Unveiling the Magic of Language

In an electronic era where connections and knowledge reign supreme, the enchanting power of language has become more apparent than ever. Its ability to stir emotions, provoke thought, and instigate transformation is truly remarkable. This extraordinary book, aptly titled "**Workbook Pages 8 1 Covalent Bonds Guided Pages Answers**," compiled by a very acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound impact on our existence. Throughout this critique, we shall delve in to the book is central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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