SYNTHESIS AND STOICHIOMETRIC ANALYSIS OF HEXAAMMINENICKEL(II)CHLORIDE

Abstract

In this experiment, we synthesize complex salt hexaamminenickel (II) chloride via crystallization and purification methods. We also add a known excess quantity of an acid to neutralize the ammonia hydroxide. This is to obtain the amount of ammonia present in the product. We start the experiment by adding the green solution which is the nickel (II) chloride with the concentrated ammonia. Then, the solution turns into purple colour. Based on the calculations, average amount of NH_2 found in the product is 33.67% while for chloride ions is 19.84%.

1.0 Introduction

Hexaamminenickel (II) chloride, $\frac{l(NH_3)_k}{l} = Cl_2$, is a coordination compound whose nickel atom and six ammonia molecules constitute the cation; the anion is the chloride ion. In aqueous solution nickel ion ($\frac{2+l}{NI^2}$: Ni(II)) is surrounded by six water molecules which are actually bonded to the central metal ion. This complex is called the hexaquanickel(II) ion. When aqueous nickel chloride is precipitated from solution, the nickel ions carry their six water molecules into the crystals and so solid nickel(II) chloride is hydrated, has the formula [$l(H_2 O)_k l Cl_2$, and is more properly called hexaquanickel(II) chloride. If ammonia is added to a solution of this salt, ammonia molecules compete with the water in bonding Ni 2+ and, because the ammonia forms a stronger bond than water, the ammonia replaces the water according to:

$$[L(H_2O)_6LCl_2 \text{ (aq)} + 6NH_3 \text{ (aq)} \rightarrow L[NH_3]_6]Cl_2 \text{ (aq)} + 6H_2O \text{ (i)}$$

Objective of this experiment is to synthesis a complex which is complex salt hexaamminenickel (II) chloride via crystallization and purification methods. Other than that, to quantify chemical component in a chemical complex via back titration method.

Synthesis And Stoichiometric Analysis Of Hexaamminenickel Ii Chloride

S Ben Porath

Synthesis And Stoichiometric Analysis Of Hexaamminenickel Ii Chloride:

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Table of Contents Synthesis And Stoichiometric Analysis Of Hexaamminenickel Ii Chloride

- 1. Understanding the eBook Synthesis And Stoichiometric Analysis Of Hexaamminenickel Ii Chloride
 - o The Rise of Digital Reading Synthesis And Stoichiometric Analysis Of Hexaamminenickel Ii Chloride
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Synthesis And Stoichiometric Analysis Of Hexaamminenickel Ii Chloride
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Synthesis And Stoichiometric Analysis Of Hexaamminenickel Ii Chloride
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Synthesis And Stoichiometric Analysis Of Hexaamminenickel Ii Chloride

- Personalized Recommendations
- Synthesis And Stoichiometric Analysis Of Hexaamminenickel Ii Chloride User Reviews and Ratings
- Synthesis And Stoichiometric Analysis Of Hexaamminenickel Ii Chloride and Bestseller Lists
- 5. Accessing Synthesis And Stoichiometric Analysis Of Hexaamminenickel Ii Chloride Free and Paid eBooks
 - Synthesis And Stoichiometric Analysis Of Hexaamminenickel Ii Chloride Public Domain eBooks
 - Synthesis And Stoichiometric Analysis Of Hexaamminenickel Ii Chloride eBook Subscription Services
 - Synthesis And Stoichiometric Analysis Of Hexaamminenickel Ii Chloride Budget-Friendly Options
- 6. Navigating Synthesis And Stoichiometric Analysis Of Hexaamminenickel Ii Chloride eBook Formats
 - o ePub, PDF, MOBI, and More
 - Synthesis And Stoichiometric Analysis Of Hexaamminenickel Ii Chloride Compatibility with Devices
 - Synthesis And Stoichiometric Analysis Of Hexaamminenickel Ii Chloride Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Synthesis And Stoichiometric Analysis Of Hexaamminenickel Ii Chloride
 - Highlighting and Note-Taking Synthesis And Stoichiometric Analysis Of Hexaamminenickel Ii Chloride
 - o Interactive Elements Synthesis And Stoichiometric Analysis Of Hexaamminenickel Ii Chloride
- 8. Staying Engaged with Synthesis And Stoichiometric Analysis Of Hexaamminenickel Ii Chloride
 - o Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - o Following Authors and Publishers Synthesis And Stoichiometric Analysis Of Hexaamminenickel Ii Chloride
- 9. Balancing eBooks and Physical Books Synthesis And Stoichiometric Analysis Of Hexaamminenickel Ii Chloride
 - Benefits of a Digital Library
 - o Creating a Diverse Reading Collection Synthesis And Stoichiometric Analysis Of Hexaamminenickel Ii Chloride
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Synthesis And Stoichiometric Analysis Of Hexaamminenickel Ii Chloride
 - Setting Reading Goals Synthesis And Stoichiometric Analysis Of Hexaamminenickel Ii Chloride
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Synthesis And Stoichiometric Analysis Of Hexaamminenickel Ii Chloride

- Fact-Checking eBook Content of Synthesis And Stoichiometric Analysis Of Hexaamminenickel Ii Chloride
- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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