

Thermochemistry

Multiple Choice Questions

1. When the three types of energy changes are arranged in order of decreasing molar heat values, what is the correct order?
 A. chemical, phase, nuclear
 B. nuclear, chemical, phase
 C. nuclear, phase, chemical
 D. phase, chemical, nuclear
2. The temperature remains constant as energy is added to a substance. How may the substance be changing?
 A. from a gas to a solid
 B. from a liquid to a gas
 C. from a liquid to a solid
 D. in the amount of kinetic energy
3. Which metal requires the most energy to raise 1.00 g of it by 1.00°C?
 A. aluminum
 B. copper
 C. lead
 D. nickel
4. A substance has a molar heat of combustion of -810.4 kJ/mol. When 0.285 mol of the substance is burned in a calorimeter containing 8.60 kg of water, what is the increase of the water temperature?
 A. 0.156°C
 B. 6.41°C
 C. 7.85°C
 D. 12.8°C
5. The addition of 9.54 kJ of heat is required to raise the temperature of 225.0 g of a liquid hydrocarbon from 20.5°C to 45.0°C. What is the heat capacity of this hydrocarbon?
 A. 0.94 J/g°C
 B. 1.73 J/g°C
 C. 1.88 J/g°C
 D. 9.42 J/g°C
6. If the molar heat of fusion of sodium is 2.63 kJ/mol, how much energy is needed to melt 180.0 g of solid sodium at its melting point?
 A. 2.63 kJ
 B. 2.93 kJ
 C. 26.6 kJ
 D. 473 kJ
7. Which process is exothermic?
 A. the boiling of liquid nitrogen
 B. the freezing of water
 C. the sublimation of dry ice
 D. the vaporization of water
8. If the heat of formation for the chemical reaction, $S_8(s) + 12O_2(g) \rightarrow 8SO_2(g)$, is -395.7 kJ/mol of SO_2 , how can the equation be rewritten to include the heat term?
 A. $S_8(s) + 12O_2(g) \rightarrow 8SO_2(g) + 3166 \text{ kJ}$
 B. $S_8(s) + 12O_2(g) \rightarrow 8SO_2(g) - 395.7 \text{ kJ}$
 C. $S_8(s) + 12O_2(g) \rightarrow 8SO_2(g) + 395.7 \text{ kJ}$
 D. $S_8(s) + 12O_2(g) + 3166 \text{ kJ} \rightarrow 8SO_2(g)$
9. Which statement correctly describes the energy changes that take place when a solid changes to a liquid at constant temperature?
 A. The potential energy increases, and the kinetic energy remains constant.
 B. The potential energy remains constant, and the kinetic energy increases.
 C. The potential energy decreases, and the kinetic energy decreases.
 D. The potential energy increases, and the kinetic energy decreases.
10. Why does warm water have a more rapid rate of evaporation than cold water?
 A. It has a higher viscosity.
 B. Its molecules have a higher average kinetic energy.
 C. Its molecules have more attraction for one another.
 D. More of its molecules have nearly the same kinetic energy.
11. If the heat of fusion of a substance is 20 kJ/mol, what heat is released when 1.0 mol of liquid at the melting point freezes?
 A. 10 kJ
 B. 20 kJ
 C. 40 kJ
 D. 80 kJ
12. The amount of 2.00 mol of a solid ionic compound was dissolved in 225 mL of water in an insulated container. The initial temperature of the water was 20.0°C and the final temperature of the water was 47.8°C. What is H_{soln} ?
 A. -27.8 kJ/mol
 B. -26.2 kJ/mol
 C. -13.1 kJ/mol
 D. -6.26 kJ/mol
13. What does it mean if the ΔH value for a chemical reaction is positive?
 A. kinetic energy is increasing in the system
 B. potential energy is decreasing in the system
 C. products have less potential energy than reactants
 D. reactants have less potential energy than products.
14. Living plants produce glucose in the process of photosynthesis according to this equation:
 $6CO_2(g) + 6H_2O(l) + \text{energy} \rightarrow C_6H_{12}O_6(s) + 6O_2(g)$
 Is this reaction endothermic or exothermic, and is the value of ΔH° positive or negative?
 A. endothermic, positive
 B. endothermic, negative
 C. exothermic, positive
 D. exothermic, negative

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Thermochemistry Mr. Ajeet Singh, 2018-05-16 Introduces thermodynamic principles related to chemical processes including enthalpy entropy Gibbs free energy and calorimetry **Internal Combustion Engines** Shyam K. Agrawal, 2006 Salient Features The New Edition Is A Thoroughly Revised Version Of The Earlier Edition And Presents A Detailed Exposition Of The Basic Principles Of Design Operation And Characteristics Of Reciprocating I C Engines And Gas Turbines Chemistry Of Combustion Engine Cooling And Lubrication Requirements Liquid And Gaseous Fuels For Ic Engines Compressors Supercharging And Exhaust Emission Its Standards And Control Thoroughly Explained Jet And Rocket Propulsion Alternate Potential Engines Including Hybrid Electric And Fuel Cell Vehicles Are Discussed In Detail Chapter On Ignition System Includes Electronic Injection Systems For Si And Ci Engines 150 Worked Out Examples Illustrate The Basic Concepts And Self Explanatory Diagrams Are Provided Throughout The Text More Than 200 Multiple Choice Questions With Answers A Good Number Of Review Questions Numerical With Answers For Practice Will Help Users In Preparing For Different Competitive Examinations With These Features The Present Text Is Going To Be An Invaluable One For Undergraduate Mechanical Engineering Students And Amie Candidates **IIT Chemistry-II** , **Chemistry for Degree Students B.Sc. (Honours) Semester II, 1/e (As per CBCS)** Madan R.L., 2022 This textbook has been designed to meet the needs of B Sc Honours Second Semester students of Chemistry as per the UGC Choice Based Credit System CBCS Maintaining the traditional approach to the subject this textbook lucidly explains the basics of Organic and Physical Chemistry Important topics such as alkanes alkenes alkynes stereochemistry aliphatic hydrocarbons thermochemistry chemical thermodynamics and chemical equilibrium are aptly discussed to give an overview of organic and physical chemistry Laboratory work has also been included to help students achieve solid conceptual understanding and learn experimental procedures **General Chemistry Workbook** Daniel C. Tofan, 2010-07-28 This workbook is a comprehensive collection of solved exercises and problems typical to AP introductory and general chemistry courses as well as blank worksheets containing further practice problems and questions It contains a total of 197 learning objectives grouped in 28 lessons and covering the vast majority of the types of problems that a student will encounter in a typical one year chemistry course It also contains a fully solved 50 question practice test which gives students a good idea of what they might expect on an actual final exam covering the entire material **Oswaal JEE Main (2019-2023) Question Bank Chapterwise + Topicwise | Chemistry (For 2024 Exam)** Oswaal Editorial Board, 2023-05-25 Description of the product 100% Updated with 4 Shifts Fully Solved 2023 January April Papers Extensive Practice No of Questions Physics 1000 Chemistry 1000 Mathematics 1000 Cognitive Learning with Smart Mind Maps Mnemonics Valuable Exam Insights with Expert Tips to crack JEE Main in first attempt Concept Clarity with Concept based revision notes detailed explanations 100% Exam Readiness with 5 Years Chapter wise Trend Analysis 2019 2023 **AP Chemistry Crash Course, Book + Online** Adrian Dingle, 2025-04-03 AP Chemistry Crash Course A Higher

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