

Practice Questions: Thermochemistry

- Calculate the work energy, w , gained or lost by the system when a gas expands from 15 L to 35 L against a constant pressure of 1.5 atm.
(1 L·atm = 100 J)
- What is expected when the reaction shown below takes place in a thermally-insulated container outfitted with a movable piston at a constant atmospheric pressure of 1 atm?
$$2\text{C}_2\text{H}_2(g) + 7\text{O}_2(g) \rightarrow 4\text{CO}_2(g) + 6\text{H}_2\text{O}(g)$$
 - Volume will decrease and work will be done on the system.
 - Volume will decrease and work will be done by the system.
 - Volume will increase and work will be done by the system.
 - Volume will decrease and work will be done on the system.
- An ideal gas expands into a vacuum (external pressure = 0) without gaining or losing heat. For this expansion...
 - ΔE increases.
 - ΔE decreases.
 - ΔE does not change.
- Which of the following is TRUE if $E_{\text{sys}} = -95 \text{ J}$?
 - The system is gaining 95 J, while the surroundings are losing 95 J.
 - The system is losing 95 J, while the surroundings are gaining 95 J.
 - Both the system and the surroundings are gaining 95 J.
 - Both the system and the surroundings are losing 95 J.
 - None of the above is true.
- Which of the following signs for q and w represent a system that is doing work on the surroundings, as well as losing heat to the surroundings?
 - $q = -$, $w = -$
 - $q = -$, $w = +$
 - $q = +$, $w = +$
 - $q = +$, $w = -$
 - None of these represent the system referenced above.
- In which case is the work on the system always positive?
 - $\Delta V < 0$
 - $\Delta E > 0$
 - $\Delta V > 0$
 - $\Delta V = 0$
 - $\Delta E = 0$
 - $\Delta E < 0$
- What is the change in internal energy for a system that absorbs 15.0 kJ of heat and performs 610 J of work on the surroundings?
- Calculate the change in internal energy (ΔE) for a system that is giving off 25.0 kJ of heat and is changing from 12.00 L to 6.00 L in volume at 1.50 atm pressure.
(Remember that 101.3 J = 1 L·atm.)
- Find ΔE° for the reaction below if the process is carried out at a constant pressure of 1.00 atm and ΔV (the volume change) equals -24.5 L .
(1 L·atm = 101 J)
$$2\text{CO}(g) + \text{O}_2(g) \rightarrow 2\text{CO}_2(g) \quad \Delta H^\circ = -566 \text{ kJ}$$
- Which of the following results in the largest decrease (most negative change) in internal energy for a system?
 - The system volume decreases and the surroundings get colder.
 - The system volume decreases and the surroundings get hotter.
 - The system volume increases and the surroundings get colder.
 - The system volume increases and the surroundings get hotter.
 - The system volume remains constant and the surroundings get colder.

Thermochemistry Practice Multiple Choice Questions

Terry C. Jones



Thermochemistry Practice Multiple Choice Questions:

Decoding **Thermochemistry Practice Multiple Choice Questions**: Revealing the Captivating Potential of Verbal Expression

In an era characterized by interconnectedness and an insatiable thirst for knowledge, the captivating potential of verbal expression has emerged as a formidable force. Its capability to evoke sentiments, stimulate introspection, and incite profound transformations is genuinely awe-inspiring. Within the pages of "**Thermochemistry Practice Multiple Choice Questions**," a mesmerizing literary creation penned by way of a celebrated wordsmith, readers set about an enlightening odyssey, unraveling the intricate significance of language and its enduring impact on our lives. In this appraisal, we shall explore the book is central themes, evaluate its distinctive writing style, and gauge its pervasive influence on the hearts and minds of its readership.

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