

## Thermochemistry Review Part 1

**Definition Questions:** you should be able to use or apply the following terms. Be able to write complete definitions for,

|                  |                        |                               |
|------------------|------------------------|-------------------------------|
| energy           | temperature            | standard state                |
| potential energy | heat                   | standard heat of formation    |
| kinetic energy   | specific heat capacity | exothermic                    |
| enthalpy         | endothermic            | Law of conservation of Energy |
| thermal          | Hess's Law             | isolated, closed, open system |

1. The specific heat capacity of diamond is  $0.5050 \text{ J/g}^\circ\text{C}$ . How much energy is required to heat  $25.0 \text{ g}$  of diamond from  $10.5^\circ\text{C}$  to  $15.6^\circ\text{C}$ ?

$$q = m \times c \times \Delta t$$

$$= 25 \text{ g} \times 0.5050 \text{ J/g}^\circ\text{C} \times (15.6 - 10.5^\circ\text{C}) = 64.39 \text{ J}$$

2. A piece of aluminum with a mass of  $50 \text{ g}$  and an initial temperature of  $90^\circ\text{C}$  is placed into  $100 \text{ mL}$  of water at a temperature of  $25^\circ\text{C}$ . The temperature of water rises to  $31.3^\circ\text{C}$ . Determine the specific heat capacity of aluminum.

$Q \text{ lost (Al)} = Q \text{ gained (Water)}$

$$-(mc\Delta t) = +mc \Delta t$$

$$-(50 \times c \times (31.3 - 90)) = (100 \times 4.18 \times (31.3 - 25))$$

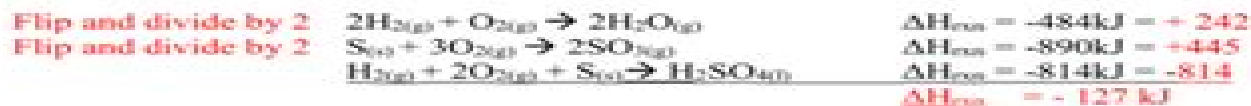
$$-(-2935 c) = 2633.4$$

$$c = 2633.4 / 2935 = 0.897 \text{ J/g}^\circ\text{C}$$

3. How much energy is absorbed by  $300 \text{ g}$  of methanol,  $\text{CH}_3\text{OH}$ , as it evaporates? Given: The molar heat of vaporization is  $35.3 \text{ kJ/mol}$ .

|                                  |                                                  |                                         |
|----------------------------------|--------------------------------------------------|-----------------------------------------|
| $q = ?$                          | $\Delta H = q/n$                                 | $n \text{ CH}_3\text{OH} = m/\text{MM}$ |
| $m = 300 \text{ g}$              | $q = H \times n$                                 | $n = 300 \text{ g} / 32 \text{ g/mol}$  |
| $\Delta H = 35.3 \text{ kJ/mol}$ | $q = (35.3 \text{ kJ} \times 9.375 \text{ mol})$ | $n = 9.375 \text{ mol}$                 |
|                                  | $q = 332.81 \text{ kJ}$                          |                                         |

4. What is the molar enthalpy of the formation of  $1 \text{ mol H}_2\text{SO}_{4(l)}$  given the following information?



# Thermochemistry Review 2 Key

**Scott C. Dulebohn**



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