#### Thermochemistry Review Part 1

**Definition Questions:** you should be able to use or apply the following terms. Be able to write complete definitions for,

energy	temperature	standard state
potential energy	heat	standard heat of formation
kinetic energy	specific heat capacity	exothermic
enthalpy	endothermic	Law of conservation of Energy
thermal	Hess's Law	isolated, closed, open system

 The specific heat capacity of diamond is 0.5050 J/g°C. How much energy is required to heat 25.0 g of diamond from 10.5°C to 15.6°C?

$$q = m \times c \times \triangle t$$
  
= 25  $g \times 0.5050 \text{ J/g}^{\circ}\text{C} \times (15.6 - 10.5^{\circ}\text{C}) = 64.39 \text{ J}$ 

 A piece of aluminum with a mass of 50g and an initial temperature of 90°C is placed into 100mL of water at a temperature of 25°C. The temperature of water rises to 31.3°C. Determine the specific heat capacity of aluminum.

```
Q lost (Al) = Q gained (Waser)

-(mc△t) = +mc △t

-(50 x c x (31.3-90)) = (100 x 4.18 x (31.3-25))

- (-2935 c) = 2633.4

c = 2633.4 / 2935 = 0.897 J/g°C
```

How much energy is absorbed by 300g of methanol, CH<sub>2</sub>OH, as it evaporates?
 Given: The molar heat of vaporization is 35.3kJ/mol.

$$\begin{array}{lll} q=? & \triangle H=q/n & n \ CH_2OH=m/MM \\ m=300 \ g & q=H \ x \ n & n=300 g \ /32 \ g/mol \\ \triangle H=35.3 \ kJ/mol & q=(35.5 \ kJ \ x \ 9.375 \ mol) & n=9.375 \ mol \\ q=332.81 \ kJ & & & \end{array}$$

What is the molar enthalpy of the formation of Imol H<sub>2</sub>SO<sub>4(l)</sub> given the following information?
 H<sub>2</sub>O<sub>1(l)</sub> + SO<sub>3(l)</sub> → H<sub>2</sub>SO<sub>4(l)</sub>
 ΔH<sub>con</sub> = xkJ

Flip and divide by 2 
$$2H_{2(g)} + O_{2(g)} \rightarrow 2H_2O_{(g)}$$
  $\Delta H_{mn} = -484kJ = +242$   
Flip and divide by 2  $S_{(s)} + 3O_{2(g)} \rightarrow 2SO_{N(g)}$   $\Delta H_{mn} = -890kJ = +445$   
 $\Delta H_{mn} = -814kJ = -814$   
 $\Delta H_{mn} = -814kJ = -814$ 

# **Thermochemistry Test Review Answers**

**ACT** 

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## **Table of Contents Thermochemistry Test Review Answers**

- 1. Understanding the eBook Thermochemistry Test Review Answers
  - The Rise of Digital Reading Thermochemistry Test Review Answers
  - Advantages of eBooks Over Traditional Books
- 2. Identifying Thermochemistry Test Review Answers
  - Exploring Different Genres
  - o Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
  - $\circ \ \ Popular \ eBook \ Platforms$
  - Features to Look for in an Thermochemistry Test Review Answers
  - User-Friendly Interface
- 4. Exploring eBook Recommendations from Thermochemistry Test Review Answers
  - Personalized Recommendations

- Thermochemistry Test Review Answers User Reviews and Ratings
- Thermochemistry Test Review Answers and Bestseller Lists
- 5. Accessing Thermochemistry Test Review Answers Free and Paid eBooks
  - Thermochemistry Test Review Answers Public Domain eBooks
  - Thermochemistry Test Review Answers eBook Subscription Services
  - Thermochemistry Test Review Answers Budget-Friendly Options
- 6. Navigating Thermochemistry Test Review Answers eBook Formats
  - o ePub, PDF, MOBI, and More
  - Thermochemistry Test Review Answers Compatibility with Devices
  - Thermochemistry Test Review Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Thermochemistry Test Review Answers
  - Highlighting and Note-Taking Thermochemistry Test Review Answers
  - Interactive Elements Thermochemistry Test Review Answers
- 8. Staying Engaged with Thermochemistry Test Review Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Thermochemistry Test Review Answers
- 9. Balancing eBooks and Physical Books Thermochemistry Test Review Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Thermochemistry Test Review Answers
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Thermochemistry Test Review Answers
  - Setting Reading Goals Thermochemistry Test Review Answers
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Thermochemistry Test Review Answers
  - Fact-Checking eBook Content of Thermochemistry Test Review Answers

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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