

Thermochemistry Practice Problems ANSWER KEY

1. Calculate the standard enthalpy of reaction for the process



$$\Delta H = -176 \text{ kJ}$$

2. Calculate the heat evolved in the combustion of ethyl alcohol, ΔH_{comb}



$$\Delta H = -1366.7 \text{ kJ}$$

3. Calculate the heat required to decompose limestone, CaCO_3 , to lime, CaO , and carbon dioxide, with all substances at standard conditions.



$$\Delta H^\circ_f \text{CaCO}_3(s) = -1206.9 \text{ kJ/mol}$$

$$\Delta H^\circ_f \text{CaO}(s) = -635.1 \text{ kJ/mol}$$

$$\Delta H^\circ_f \text{CO}_2(g) = -393.5 \text{ kJ/mol}$$

$$\Delta H = 178.3 \text{ kJ}$$

4. $\text{CH}_4(g) + 3 \text{Cl}_2(g) \rightarrow \text{CHCl}_3(l) + 3 \text{HCl}(g) \quad \Delta H = ?$



$$\Delta H = -305.2 \text{ kJ}$$

5. Cyanamide, CH_2N_2 , is a weak acid that is sometimes used as a fertilizer. Calculate the standard enthalpy of formation for cyanamide, given the following standard enthalpies of reaction:



$$\Delta H = 62.1 \text{ kJ}$$

6. The standard molar enthalpy change is -905.2 kJ for the oxidation of ammonia.



Calculate the standard molar enthalpy of formation for ammonia based on the following standard enthalpies of formation: $\Delta H^\circ_f \text{NO}(g) = 90.3 \text{ kJ/mol}$ and $\Delta H^\circ_f \text{H}_2\text{O}(g) = -241.8 \text{ kJ/mol}$

$$-46.1 \text{ kJ/mol}$$

7. The thermite reaction: $2 \text{Al}(s) + \text{Fe}_2\text{O}_3(s) \rightarrow \text{Al}_2\text{O}_3(s) + 2 \text{Fe}(s) + 852 \text{ kJ}$

If you begin with 10.0 g of Al and excess Fe_2O_3 , how many kJ of heat are evolved at constant pressure?

$$158 \text{ kJ}$$

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Chemistry: 1001 Practice Problems For Dummies (+ Free Online Practice) Heather Hattori, Richard H.

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Patrick E. McMahon, Rosemary McMahon, Bohdan Khomtchouk, 2019-02-13 This work evolved over thirty combined years of teaching general chemistry to a variety of student demographics The focus is not to recap or review the theoretical concepts well described in the available texts Instead the topics and descriptions in this book make available specific detailed step by step methods and procedures for solving the major types of problems in general chemistry Explanations instructional process sequences solved examples and completely solved practice problems are greatly expanded containing significantly more detail than can usually be devoted to in a comprehensive text Many chapters also provide alternative viewpoints as an aid to understanding Key Features The authors have included every major topic in the first semester of general chemistry and most major topics from the second semester Each is written in a specific and detailed step by step process for problem solving whether mathematical or conceptual Each topic has greatly expanded examples and solved practice problems containing significantly more detail than found in comprehensive texts Includes a chapter designed to eliminate confusion concerning acid base reactions which often persists through working with acid base equilibrium Many chapters provide alternative viewpoints as an aid to understanding This book addresses a very real need for a large number of incoming freshman in STEM fields *Chemistry Workbook For Dummies with Online Practice* Chris Hren, Peter J. Mikulecky, 2017-03-21 Take the confusion out of chemistry with hundreds of practice problems Chemistry Workbook For Dummies is your ultimate companion for introductory chemistry at the high school or college level Packed with hundreds of practice problems this workbook gives you the practice you need to internalize the essential concepts that form the foundations of chemistry From

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Cracking the GRE Chemistry Subject Test Princeton Review (Firm), 2005 The GRE subject tests are among the most difficult standardized exams Rather than testing general problem solving skills they require highly specialized knowledge The experts at The Princeton Review have thoroughly researched each subject test to provide students with the most thorough up to date information available Students don't need to relearn the entire histories of their fields just what they need to know to earn high scores on the exams Each guide includes one full length practice exam complete with comprehensive explanations for every solution

General Chemistry Donald A. McQuarrie, Stanley Gill, 2011-06-15 This Fourth Edition of McQuarrie's classic text offers a thorough revision and a quantum leap forward from the previous edition Taking an atoms first approach it promises to be another ground breaking text in the tradition of McQuarrie's many previous works This outstanding new text available in a soft cover edition offers professors a fresh choice and outstanding value

Elementary Chemical Reactor Analysis Rutherford Aris, 2013-09-03 Elementary Chemical Reactor Analysis focuses on the processes reactions methodologies and approaches involved in chemical reactor analysis including stoichiometry

adiabatic reactors external mass transfer and thermochemistry The publication first takes a look at stoichiometry and thermochemistry and chemical equilibrium Topics include heat of formation and reaction measurement of quantity and its change by reaction concentration changes with a single reaction rate of generation of heat by reaction and equilibrium of simultaneous and heterogeneous reactions The manuscript then offers information on reaction rates and the progress of reaction in time Discussions focus on systems of first order reactions concurrent reactions of low order general irreversible reaction variation of reaction rate with extent and temperature and heterogeneous reaction rate expressions The book examines the interaction of chemical and physical rate processes continuous flow stirred tank reactor and adiabatic reactors Concerns include multistage adiabatic reactors adiabatic stirred tank stability and control of the steady state mixing in the reactor effective reaction rate expressions and external mass transfer The publication is a dependable reference for readers interested in chemical reactor analysis

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