

Thermochemistry Practice Problems

KEY

Flow of Energy:

1. How many kJ of energy are in a donut that contains 200.0 Calories?

$$\frac{200 \text{ Cal}}{1 \text{ Cal}} \times \frac{4.184 \text{ J}}{1 \text{ Cal}} = \boxed{836.8 \text{ kJ}}$$

2. What is the specific heat of a substance that has a mass of 25.0 g and requires 525.0 calories to raise its temperature by 15.0°C?

$$q = mc\Delta T$$

$$m = 25.0 \text{ g}$$

$$q = 525.0 \text{ cal}$$

$$\Delta T = 15.0^\circ\text{C}$$

$$525.0 \text{ cal} = (25.0 \text{ g})(c)(15.0^\circ\text{C})$$

$$\boxed{c = 1.40 \text{ cal/g}^\circ\text{C}}$$

3. Suppose 100.0 g of H₂O(s) absorbs 1255.0 J of heat. What is the corresponding temperature change?

The specific heat capacity of H₂O(s) is 2.1 J/g°C. $q = mc\Delta T$

$$q = 1255.0 \text{ J}$$

$$m = 100.0 \text{ g}$$

$$c = 2.1 \text{ J/g}^\circ\text{C}$$

$$1255.0 \text{ J} = (100.0 \text{ g})(2.1 \text{ J/g}^\circ\text{C})(\Delta T)$$

$$\boxed{\Delta T = 6.0^\circ\text{C}}$$

4. How many joules of heat energy are required to raise the temperature of 100.0 g of aluminum by 120.0°C? The specific heat capacity of aluminum is 0.90 J/g°C.

$$q = mc\Delta T$$

$$m = 100.0 \text{ g}$$

$$c = 0.90 \text{ J/g}^\circ\text{C}$$

$$\Delta T = 120.0^\circ\text{C}$$

$$q = (100.0 \text{ g})(0.90 \text{ J/g}^\circ\text{C})(120.0^\circ\text{C})$$

$$\boxed{q = 11,000 \text{ J}}$$

5. In reference to #4, what is the molar heat capacity of aluminum?

$$(0.90 \text{ J/g}^\circ\text{C})(26.98 \text{ g/mol}) = \boxed{24 \text{ J/mol}^\circ\text{C}}$$

Measuring and Expressing Heat Changes:

6. A student mixed 75.0 mL of water containing 0.75 mol HCl at 25°C with 75.0 mL of water containing 0.75 mol of NaOH at 25°C in a foam cup calorimeter. The temperature of the resulting solution increased to 35°C. How much heat in kilojoules was released by this reaction ($C_{\text{water}} = 4.18 \text{ J/g}^\circ\text{C}$)?

$$q = mc\Delta T$$

$$m = 150.0 \text{ g}$$

$$c = 4.184 \text{ J/g}^\circ\text{C}$$

$$\Delta T = 10^\circ\text{C}$$

$$q = (150.0 \text{ g})(4.184 \text{ J/g}^\circ\text{C})(10^\circ\text{C})$$

$$\boxed{q = 6300 \text{ J}}$$

7. Calculate the amount of heat evolved when 150 g of water are reacted with excess calcium according to the following thermochemical equation:



$$\frac{150 \text{ g H}_2\text{O}}{18.02 \text{ g H}_2\text{O}} \times \frac{1 \text{ mol H}_2\text{O}}{2 \text{ mol H}_2\text{O}} \times \frac{-65.2 \text{ kJ}}{1} = \boxed{-270 \text{ kJ}; \text{ exothermic}}$$

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Khomtchouk, 2019-02-13 This work evolved over thirty combined years of teaching general chemistry to a variety of student demographics The focus is not to recap or review the theoretical concepts well described in the available texts Instead the topics and descriptions in this book make available specific detailed step by step methods and procedures for solving the major types of problems in general chemistry Explanations instructional process sequences solved examples and completely solved practice problems are greatly expanded containing significantly more detail than can usually be devoted to in a comprehensive text Many chapters also provide alternative viewpoints as an aid to understanding Key Features The authors have included every major topic in the first semester of general chemistry and most major topics from the second semester Each is written in a specific and detailed step by step process for problem solving whether mathematical or conceptual Each topic has greatly expanded examples and solved practice problems containing significantly more detail than found in comprehensive texts Includes a chapter designed to eliminate confusion concerning acid base reactions which often persists through working with acid base equilibrium Many chapters provide alternative viewpoints as an aid to understanding This book addresses a very real need for a large number of incoming freshman in STEM fields

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Table of Contents Thermochemistry Practice Problems Answer Ch 6

1. Understanding the eBook Thermochemistry Practice Problems Answer Ch 6
 - The Rise of Digital Reading Thermochemistry Practice Problems Answer Ch 6
 - Advantages of eBooks Over Traditional Books
2. Identifying Thermochemistry Practice Problems Answer Ch 6
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Thermochemistry Practice Problems Answer Ch 6
 - User-Friendly Interface
4. Exploring eBook Recommendations from Thermochemistry Practice Problems Answer Ch 6
 - Personalized Recommendations
 - Thermochemistry Practice Problems Answer Ch 6 User Reviews and Ratings

- Thermochemistry Practice Problems Answer Ch 6 and Bestseller Lists
- 5. Accessing Thermochemistry Practice Problems Answer Ch 6 Free and Paid eBooks
 - Thermochemistry Practice Problems Answer Ch 6 Public Domain eBooks
 - Thermochemistry Practice Problems Answer Ch 6 eBook Subscription Services
 - Thermochemistry Practice Problems Answer Ch 6 Budget-Friendly Options
- 6. Navigating Thermochemistry Practice Problems Answer Ch 6 eBook Formats
 - ePub, PDF, MOBI, and More
 - Thermochemistry Practice Problems Answer Ch 6 Compatibility with Devices
 - Thermochemistry Practice Problems Answer Ch 6 Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Thermochemistry Practice Problems Answer Ch 6
 - Highlighting and Note-Taking Thermochemistry Practice Problems Answer Ch 6
 - Interactive Elements Thermochemistry Practice Problems Answer Ch 6
- 8. Staying Engaged with Thermochemistry Practice Problems Answer Ch 6
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Thermochemistry Practice Problems Answer Ch 6
- 9. Balancing eBooks and Physical Books Thermochemistry Practice Problems Answer Ch 6
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Thermochemistry Practice Problems Answer Ch 6
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Thermochemistry Practice Problems Answer Ch 6
 - Setting Reading Goals Thermochemistry Practice Problems Answer Ch 6
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Thermochemistry Practice Problems Answer Ch 6
 - Fact-Checking eBook Content of Thermochemistry Practice Problems Answer Ch 6
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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